

# Fractional crystallization of primitive, hydrous arc magmas: an experimental study at 0.7 GPa

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**Abstract** Differentiation of mantle-derived, hydrous, basaltic magmas is a fundamental process to produce evolved intermediate to SiO<sub>2</sub>-rich magmas that form the bulk of the middle to shallow continental and island arc crust. This study reports the results of fractional crystallization experiments conducted in a piston cylinder apparatus at 0.7 GPa for hydrous, calc-alkaline to arc tholeiitic magmas. Fractional crystallization was approached by synthesis of starting materials representing the liquid composition of the previous, higher temperature experiment. Temperatures ranged from near-liquidus at 1,170 °C to near-solidus conditions at 700 °C. H<sub>2</sub>O contents varied from 3.0 to more than 10 wt%. The liquid line of descent covers the entire compositional range from olivine–tholeiite (1,170 °C) to high-silica rhyolite (700 °C) and evolves from metaluminous to peraluminous compositions. The following crystallization sequence has been established: olivine → clinopyroxene → plagioclase, spinel → orthopyroxene, amphibole, titanomagnetite → apatite → quartz, biotite. Anorthite-rich plagioclase and spinel are responsible for a marked increase in SiO<sub>2</sub>-content (from 51 to 53 wt%) at 1,040 °C. At lower temperatures,

fractionation of amphibole, plagioclase and Fe–Ti oxide over a temperature interval of 280 °C drives the SiO<sub>2</sub> content continuously from 53 to 78 wt%. Largest crystallization steps were recorded around 1,040 °C and at 700 °C. About 40 % of ultramafic plutonic rocks have to crystallize to generate basaltic–andesitic liquids, and an additional 40 % of amphibole–gabbroic cumulate to produce granitic melts. Andesitic liquids with a liquidus temperature of 1,010 °C only crystallize 50 % over an 280 °C wide range to 730 °C implying that such liquids form mobile crystal mushes (<50 % crystals) in long-lived magmatic systems in the middle crust, allowing for extensive fractionation, assimilation and hybridization with periodic replenishment of more mafic magmas from deeper magma reservoirs.

**Keywords** Liquid line of descent · Fractional crystallization · Calc-alkaline magmas · Mid-crustal magma reservoirs

## Introduction

Voluminous intermediate to granitic rocks forming large batholiths and associated volcanic provinces typically occur in supra-subduction, arc-related settings such as the Peruvian coastal batholiths, the Californian batholith (USA), the Patagonian batholith (Chile), the Himalaya batholith or on a smaller scale, the Tertiary Adamello batholith (Northern Italy). For their origin, a number of processes have been invoked such as fractional crystallization from mantle-derived basaltic precursors (e.g., Bowen 1928; Cawthorn et al. 1976; Sisson and Grove 1993a, b; Ulmer 2001), additional extensive assimilation of and hybridization with crustally derived partial melts

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generated by heating of mantle-derived magmas over prolonged time (Hildreth and Moorbath 1988; Blatter et al. 2013) as well as exclusive partial melting of deep-crustal lithologies (mafic amphibolites to granulites, e.g., Rapp and Watson (1995), as well as pelitic micaschists, e.g., Clemens et al. (2011)). In order to identify and quantify the relative significance of these processes, phase equilibria, major and trace element compositions, and partitioning behavior of the products generated by the proposed end-member processes need to be known. In this contribution, we focus on fractional crystallization of hydrous, calc-alkaline to tholeiitic arc-type primary to parental melts at pressure conditions corresponding to mid-crustal levels (0.7 GPa, 20–25 km). Our data complement similar studies conducted in our laboratory and elsewhere that cover the deep-crustal levels (e.g., Alonso-Perez 2006; Whitaker et al. 2008; Ulmer et al. 2014). The focus of this study is twofold: (1) To identify the phase equilibrium controls on the fractionation paths of calc-alkaline magmas and (2) to quantify the liquid line of descent from the basaltic liquidus to granitic derivative liquids close to the granite solidus. This study contributes at evaluating the mineralogy, mineral chemistry and modal proportions of products generated through fractional crystallization and ultimately provides arguments in favor of, or in contradiction to the potential derivation of silicic magmas by fractional crystallization of precursor basaltic magmas, a subject that is highly controversial (e.g., Clemens et al. 2011; Jagoutz et al. 2011; Melekhova et al. 2013) but intimately linked to the generation of the continental crust.

An ancillary goal of these experimental series at 0.7 GPa was to test the hypothesis that up to 50 % crystallization of ultramafic rocks from initial, mantle-derived liquids must have taken place to produce andesitic liquids

(e.g., Kay and Kay 1991; Müntener et al. 2001; Müntener and Ulmer 2006; Jagoutz et al. 2011). Quantitative assessment of the compositions of derivative liquids as well as the modal proportions of fractionated solids furthermore directly contributes to a better understanding of the rheology of such magmas as crystallinity and rheology of suspending silicate liquid are the two most important parameters governing the ascent and intrusion/eruption modes of intermediate to silicic magmas (e.g., Clemens and Mawer 1992; Brown and Solar 1998; Paterson et al. 1998; Caricchi et al. 2007).

## Experimental techniques

### Experimental strategy

In this study, (near) fractional crystallization was approached in a stepwise manner. A synthetic equivalent of a natural, near-primary olivine–tholeiite composition was used as initial starting material containing 3.0 wt% H<sub>2</sub>O to locate the liquidus (Table 1). Thereafter, the liquid composition coexisting with small to moderate amounts of solid phases was determined by electron probe microanalysis (EPMA), and for any subsequent experiment, a new starting material corresponding to the liquid composition of the previous, higher temperature experiment was synthesized. The temperature steps between successive experiments were generally 30 °C; at lowest temperatures, 50 °C steps were employed and for intervals where significant (>20 %) crystallization occurred an additional experiment was set between 2 fractionation steps. All experiments were evaluated by least-square regression analysis using the known starting composition and the

**Table 1** Compositions of starting materials in wt%

St.mat.	SiO <sub>2</sub>	TiO <sub>2</sub>	Al <sub>2</sub> O <sub>3</sub>	Fe <sub>2</sub> O <sub>3</sub>	FeO	MnO	MgO	CaO	Na <sub>2</sub> O	K <sub>2</sub> O	P <sub>2</sub> O <sub>5</sub>	H <sub>2</sub> O	Total
RDC156	48.43	0.68	15.19	1.64	6.65	0.15	10.13	11.57	1.86	0.59	0.10	3.00	100.00
RN4	48.42	0.72	15.42	1.72	6.80	0.17	9.39	11.71	1.90	0.61	0.10	3.04	100.00
RN6V2	48.53	0.74	15.97	1.73	6.78	0.17	8.46	11.75	1.97	0.63	0.11	3.17	100.00
RN7	48.34	0.78	18.01	1.74	6.80	0.17	6.95	10.42	2.26	0.73	0.12	3.69	100.00
RN8	50.20	0.84	18.82	1.71	6.71	0.17	4.54	8.02	2.90	0.97	0.16	4.96	100.00
RN9	53.85	0.63	18.58	1.31	5.16	0.17	2.64	6.78	3.21	1.20	0.21	6.26	100.00
RN10	56.05	0.60	18.18	0.90	3.26	0.16	2.66	6.45	3.40	1.29	0.22	6.82	100.00
RN11	57.92	0.61	16.72	0.88	3.04	0.18	2.28	5.22	3.63	1.46	0.26	7.81	100.00
RN12	59.62	0.50	16.62	0.67	2.20	0.17	1.50	4.57	3.81	1.58	0.28	8.49	100.00
RN13V2	61.40	0.26	15.67	0.47	1.47	0.18	1.59	3.79	4.00	1.72	0.30	9.15	100.00
RN14V2	61.49	0.25	15.28	0.46	1.38	0.17	1.30	3.26	4.06	1.76	0.23	10.36	100.00
RN15V2	63.83	0.17	14.71	0.34	0.96	0.17	1.26	2.49	4.29	1.90	0.16	9.70	100.00
RN16V2	65.60	0.12	13.52	0.26	0.66	0.15	0.90	1.34	4.45	2.11	0.10	10.79	100.00
RN17V2	67.99	0.11	12.24	0.20	0.45	0.11	0.64	0.52	4.44	2.37	0.04	10.89	100.00

**Table 2** Experimental run conditions and phase relations of fractional crystallization experiments at 0.7 GPa (inner/single capsules)

Run #	sm <sup>a</sup>	Temp (°C)	Time (h)	Capsule material	Phase assemblages	Phase proportions <sup>b</sup> (wt%)	$\sum R^2$	$\Delta Fe$ (%rel.)	fO <sub>2</sub> $\Delta NNO$	Liq (abs)
RN1	RDC156	1,170	4	Au <sub>90</sub> Pd <sub>10</sub>	liq, v	100	0.52	-1.4		100.0
RN2	RDC156	1,150	28	Au <sub>90</sub> Pd <sub>10</sub>	liq, ol, v	98.1(8), 1.9(6)	0.37	-1.5	0.33	98.1
RN5	RN4	1,100	12	Au <sub>90</sub> Pd <sub>10</sub>	liq, ol, cpx, v	94.1(13), 3.8(5), 2.1(14)	0.20	-2.4	1.07	92.3
RN6	RN6V2	1,070	24	Au <sub>90</sub> Pd <sub>10</sub>	liq, ol, cpx, v	86.1(6), 0.9(3), 13.0(6)	0.06	9.1		79.5
RN7W2	RN7	1,055	47	Au <sub>90</sub> Pd <sub>10</sub>	liq, cpx, v	94.7(27), 5.3(4)	0.03	6.5		75.3
RN7	RN7	1,040	48	Au <sub>90</sub> Pd <sub>10</sub>	liq, cpx, (opx), plag, sp, v	75.7(43), 19.7(29), 2.2(36), 0.72 <sup>c</sup>	0.80	6.5		60.2
RN8W2	RN8	1,025	49	Au <sub>90</sub> Pd <sub>10</sub>	liq, (plag), v	100	0.11	5.5		60.2
RN8	RN8	1,010	48	Au <sub>90</sub> Pd <sub>10</sub>	liq, opx, plag, amph, mag, v	74.4(7), 0.8(10), 6.8(11), 16.6(16), 1.4(5)	0.01	2.4		44.8
RN9	RN9	980	48	Au <sub>90</sub> Pd <sub>10</sub>	liq, plag, mag, v	91.8(11), 5.1(12), 0.7 <sup>c</sup>	0.28	-0.4		41.1
RN10s	RN10	950	48	Au	liq, opx, plag, amph, v	86.0(16), 0.4(29), 10.0(24), 3.6(48)	0.18	5.0		35.3
RN11V2	RN11	920	56	Au	liq, plag, amph, mag, (usp), v	90.9(7), 2.8(6), 5.5(8), 0.18 <sup>c</sup>	0.06	-0.2		32.1
RN12V2-2	RN12	890	56	Au	liq, plag, amph, mag, v	86.4(11), 10.9(11), 1.3(14), 0.18 <sup>c</sup>	0.21	2.3		27.8
RN13V2	RN13V2	860	72	Au	liq, plag, amph, ap, v	96.1(19), 2.0(16), 1.9(23), 0.16 <sup>c</sup>	0.70	-0.6		26.7
RN14V2	RN14V2	830	75	Au	liq, plag, amph, mag, ap, v	93.0(27), 5.6(22), 0.7(34), 0.8(7), 0.20 <sup>c</sup>	1.04	1.8		24.8
RN15V2	RN15V2	780	100	Au	liq, plag, amph, mag, ap, v	88.5(20), 8.7(18), 2.6(22), 0.3(5), 0.18 <sup>c</sup>	0.54	1.7		22.0
RN16V2	RN16V2	730	100	Au	liq, plag, amph, mag, ap, v	87.8(25), 10.4(22), 1.7(20), 0.2(5), 0.07 <sup>c</sup>	0.52	1.6		19.3
RN17V2	RN17V2	700	168	Au	liq, plag, mag, qtz, bt, ap, v	73.9(20), 12.0(8), 0.2(1), 10.5(10), 3.4(4), 0.03	0.55	-0.9		14.2

Phases in brackets correspond to phases present in experiment, but whose modal proportions were insignificant (not calculated)

Abbreviations for phases: liq—liquid (glass), ol—olivine, cpx—high-Ca clinopyroxene, opx—low-Ca orthopyroxene, plag—plagioclase, sp—hercynitic spinel, mag—titanomagnetite, usp—ulvospinel, amph—amphibole, ap—apatite, qtz—quartz, bt—biotite, v—vapor (bubble)

$\Delta Fe$ : FeO loss or gain (in relative percent; negative numbers = gain) from the starting material based on mass balance calculations

$\sum R^2$ : residual sum of squares

fO<sub>2</sub>  $\Delta NNO$ :  $\log fO_2(NNO) - \log fO_2(\text{experiment})$  based on ol-liq Fe<sup>2+</sup>-Mg partitioning and the model of Kress and Carmichael (1988)

<sup>a</sup> Compositions of starting materials are given in Table 1

<sup>b</sup> Units in parentheses represent the error from regression analysis. Accordingly, 94.1 (13) is read as  $94.1 \pm 1.3$

<sup>c</sup> Phase proportion of Fe-Ti oxide determined by image analysis

compositions of all coexisting phases (glass and solid phase) to compute modal proportions (Tables 1, 2) to ensure closed system behavior, to detect potential disequilibrium, to calculate the correct amount of H<sub>2</sub>O to add to the next, lower-temperature starting material, and to account for potential alkali deficits in the EPMA analyses (mainly caused by Na migration in silicate liquids), and iron loss to noble-metal sample containers.

#### Starting materials

A synthetic equivalent of a near-primary olivine-tholeiite dike composition from the southern part of the Tertiary

Adamello batholith (RC156, coordinates: 45°57'08.04"N 10°26'33.08"E elevation: 2,344 m) was used as an initial starting composition (Table 1). The particular dike consists of olivine (forsterite content (Fo) 91.1) and clinopyroxene (Mg-number (Mg#) = molar MgO/MgO + FeO<sub>tot</sub> of 0.87) phenocrysts both containing minute, idiomorphic Cr-spinel inclusions in a fine-grained matrix composed of plagioclase, amphibole and minor magnetite (Ulmer 1986). The trace element signature clearly reveals the subduction-related provenance (HFSE depletion, LREE and LILE enrichment) of this dike that is post-plutonic, cross-cutting the tonalites of the southern Adamello batholiths (Hürli-mann et al. 2014). A similar, but even more primitive

composition (RC158c), a microbasalt from the same dike–rock association has been utilized as starting material for the higher pressure, equilibrium and fractional crystallization experiments at 1.0 (Müntener and Ulmer 2006; Ulmer et al. 2014) and 1.5 GPa (Alonso-Perez 2006; Müntener and Ulmer 2006). In order to avoid a large number of experiments that solely crystallize olivine, a less Mg-rich composition has been selected for this study. However, the two compositions, RC158c and RC156, are part of the same dike generation and are directly linked by moderate (up to 15 %) olivine fractionation (see Hürlimann et al. 2014).

All liquid compositions (Table 1) were mixed from fired (1,000 °C: SiO<sub>2</sub>, TiO<sub>2</sub>, Fe<sub>2</sub>O<sub>3</sub>, MgO and CaSiO<sub>3</sub>) and dried (220–350 °C:  $\gamma$ -Al<sub>2</sub>O<sub>3</sub>, MnO, Fe<sub>2</sub>SiO<sub>4</sub>, Na<sub>2</sub>SiO<sub>3</sub>, K<sub>4</sub>Si<sub>5</sub>O<sub>12</sub> and Ca<sub>5</sub>(PO<sub>4</sub>)<sub>3</sub>(OH) oxides and silicates), and dried mixtures of AlOOH (200 °C) and Al(OH)<sub>3</sub> (100 °C) to adjust the desired H<sub>2</sub>O content provided the alumina concentration of the target composition was high enough. In the last fractionation steps where Al<sub>2</sub>O<sub>3</sub> concentrations were low and H<sub>2</sub>O concentrations high, distilled water was added by a micro-syringe to an anhydrous powder mixture during preparation of the capsules. In order to determine trace element partitioning between solid phases and the quenched liquids (glass), all starting material additionally contained 2 wt% of a trace element-doped diopside glass (32 trace element at 2,000  $\mu$ g/g level, except Th and U (1,000  $\mu$ g/g) providing 40 (U, Th 20)  $\mu$ g/g of trace element encompassing all important classes of trace elements (LILE, REE, HFSE, transition metals). The thoroughly mixed, fine-grained oxide–silicate–hydrate starting materials were directly employed in this experimental study without any further treatment.

Oxygen fugacity (fO<sub>2</sub>) was constrained by fixing the ratio of ferrous to ferric iron in the starting material using the algorithm of Kress and Carmichael (1991) for liquids to a value corresponding to the Ni–NiO buffer equilibrium (NNO) at the conditions of the experiment (pressure of 0.7 GPa and respective run temperature). This approach is only strictly valid when the charge would be composed entirely of liquid; therefore, in the present work, this is only an approximation as the amount of liquid varied between 100 and 75 %. In the absence of Fe–Ti oxides, the fO<sub>2</sub> most likely increases in the charges due to the more incompatible behavior of ferric relative to ferrous iron in the presence of olivine and clinopyroxene that dominate in the higher temperature steps. At lower temperature, prediction is more difficult as both Fe–Ti oxide and amphibole can accommodate significant amounts of ferric relative to ferrous iron. However, in these experiments, the liquid fraction always exceeded 84 %; therefore, we do not expect significant deviation from the target fO<sub>2</sub> due to crystallization of the solid phases. However, a prerequisite for successful application of this technique is the absence

of significant hydrogen and/or iron loss during the experiment, i.e., the system behaved as a closed system. Hydrogen diffusion was minimized by employing the double-capsule technique described by Kägi et al. (2005). Fe loss was minimized using Au<sub>90</sub>Pd<sub>10</sub> sample container material above 1,000 °C and pure Au below 1,000 °C and by minimizing the run duration to the minimum required for equilibration.

In runs where olivine formed a liquidus phase (at highest temperatures), the olivine–liquid Fe<sup>2+</sup>–Mg partitioning could be utilized to estimate fO<sub>2</sub> during the experiments (Ulmer 1989; Kägi et al. 2005) (see Table 2) and confirmed that fO<sub>2</sub> remained close to slightly above the targeted NNO values.

**Table 3** H<sub>2</sub>O contents of liquids (glasses) in fractional crystallization experiments at 0.7 GPa calculated from added H<sub>2</sub>O and determined by Raman spectroscopy

Run#	Temp. (°C)	Capsule	H <sub>2</sub> O added (wt%) <sup>a</sup>	H <sub>2</sub> O Raman (wt%) <sup>b</sup>	2 $\sigma$ <sup>c</sup>
RN1	1,170	inner	3.00	3.6	0.2
		outer	3.06	3.3	0.1
RN2	1,140	inner	3.06	3.1	0.1
		outer	3.06	3.3	0.1
RN4	1,120	inner	3.05	5.1	0.1
		outer	3.05	3.9	0.1
RN5	1,110	inner	3.23	3.2	0.1
		outer	3.17	3.4	0.2
RN6	1,070	inner	3.68	4.2	0.2
		outer	3.77	4.4	0.4
RN7	1,040	inner	4.87	5.3	0.1
		outer	4.84	5.0	0.2
RN8	1,010	inner	6.20	6.2	0.1
		outer	6.06	6.8	0.5
RN10s	950	single	7.84	7.8	0.4
RN11	920	single	8.46	8.7	0.3
RN12V2-2	890	single	9.79	7.9	0.3
RN13V2	860	single	9.48	9.9	0.4
RN14V2	830	single	11.12	9.6	0.3
RN15V2	780	single	10.90	9.8	0.6
RN16V2	730	single	12.25	10.7	0.9
RN17V2	700	single	12.88	n.d.	

n.d. not determined—the composition of RN17V2 is outside the calibration range

<sup>a</sup> H<sub>2</sub>O added: H<sub>2</sub>O in the liquid phase calculated from the H<sub>2</sub>O added to the starting material as Al hydroxide or by micro-syringe taking into account modal proportions of liquid, amphibole and biotite

<sup>b</sup> H<sub>2</sub>O determined by Raman spectroscopy, the error associated with the calibration (standard) and fitting of the equation used to calculate the H<sub>2</sub>O contents amounts to approximately 0.5 wt%

<sup>c</sup> Standard error (2 $\sigma$ ) of the Raman spectroscopic measurements: 9 measurements (3 analyses per point, 3 different points) were averaged for every experiment listed

## Experimental setup

The fine-grained oxide, silicate, hydroxide starting material was dried at 110 °C before packed into the noble-metal containers. At temperatures exceeding 1,000 °C, Au–Pd alloys (90:10 ratio and 80:20 for some outer capsules) were employed as capsule material combined with the capsule-in-capsule technique (CICT; Kägi et al. 2005) (Electronic Supplementary Material ESM 1). Below 1,000 °C, a single Au capsule was used assuming Fe loss to be less significant. The CICT resulted in negligible iron loss in the mass balance calculations, and Fe concentrations in the inner capsule walls were below the detection limit of the EPMA. The case of single Au capsules was somewhat different (e.g., Ratajeski and Sisson 1999): EPMA analyses of capsule walls resulted in very low iron contents and mass balance calculations were inappropriate as a Fe–Ti oxide was always present. In this case, mass balance cannot be employed to evaluate potential iron loss, as this would simply be masked by increased calculated Fe–Ti oxide modal concentrations. FeO and MgO concentrations are rather low (MgO: 0.13–2.32 wt%; FeO<sub>tot</sub>: 0.30–4.22 wt%) in these runs. The relative Fe loss has been up to 5 % and resulted in somewhat increased Mg# at lower temperatures as MgO concentrations were overestimated relative to FeO<sub>tot</sub> concentrations (e.g., Ratajeski and Sisson 1999).

After tightly filling the capsules were weight, crimped, cut and welded shut using an arc welder. Sealing of the capsules was checked by weighting before and after welding by submersion in acetone. The capsules were embedded vertically in MgO–ceramic and contained in a NaCl–pyrex–graphite–MgO piston cylinder assembly. The capsule was placed in the hotspot of the assembly, which was computed using the numerical model of Hernlund et al. (2006). All experiments were performed in a 14-mm bore end-loaded piston cylinder apparatus. A friction correction of –3 % was applied to the nominal pressure, based on calibrations employing the fayalite + qtz = orthoferrosilite univariant reaction at 1,000 °C and 1.41 GPa (Bohlen et al. 1980) and the qtz = coesite transition at 1,000 °C and 3.07 GPa (Bose and Ganguly 1995). The temperature was measured with a B-type Pt<sub>94</sub>Rh<sub>6</sub>/Pt<sub>70</sub>Rh<sub>30</sub> thermocouple, with an estimated accuracy of ±10 °C. No corrections for the pressure effects on the electromotive force (emf) were performed. All runs were first pressurized to 0.4 GPa at room temperature, then temperature increased to 550 °C and subsequently both pressure and temperature were increased concomitantly to the respective run conditions. The experiments were terminated by switching off the power supply resulting in quenching rates in excess of 100 °C/s.

## Analytical methods

All recovered charges were embedded in epoxy resin, ground to expose a longitudinal section of the capsule(s) and polished with diamond paste down to 1 micrometer. For EPMA and SEM analyses, samples were coated with 20 nm of carbon.

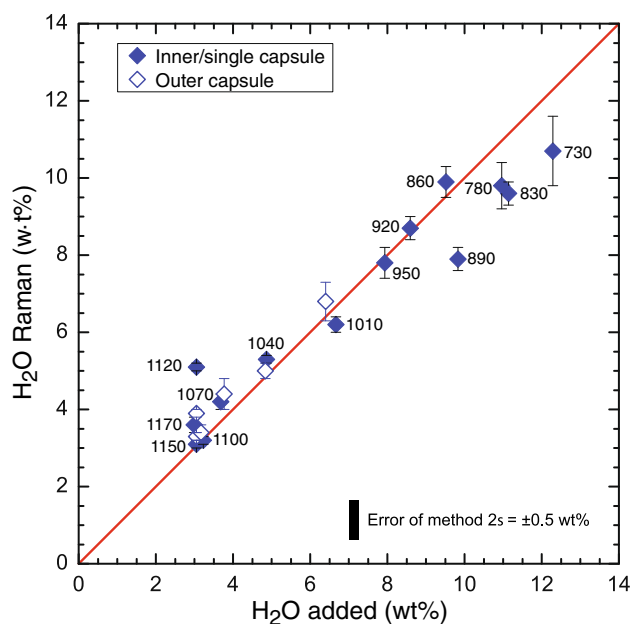
### *Electron probe micro-analyzer (EPMA)*

Experimental run products were analyzed with wavelength-dispersive spectrometers (WDS) on a JEOL JXA-8200 electron probe micro-analyzer. An acceleration voltage of 15 kV and a variable beam current and beam size (7 nA and 10–20 μm for glasses and 20 nA and 1 μm for minerals) were used to minimize alkali migration during analyses of hydrous glasses. Peak and background counting times were 20 s. All data were corrected with the PRZ model (Goldstein et al. 1992; Korolyuk et al. 2009). For glasses, H<sub>2</sub>O was used as additional element for PRZ corrections and calculated as difference to 100 wt%.

Modal proportions were calculated using non-weighted, least-squares regression (LSR) analyses implemented in Microsoft EXCEL<sup>®</sup> 2010 balancing the nominal composition of the starting material against the averages of all analyzed phases in the experimental charges. Because Fe–Ti oxide phases are overestimated using the LSR analyses in case of Fe loss to the noble-metal containers, back-scattered electron (BSE) images of entire charges and an image processing program (ImageJ) were used to determine the modal proportions of iron oxide phases (Table 2).

### *Micro-Raman spectroscopy*

The H<sub>2</sub>O contents of the recovered experimental glasses were quantified by micro-Raman spectroscopy. Raman scattering was excited using a 514.5-nm Ar Laser (Coherent Innova 90C), and measurements were performed with a Dilor Labram II confocal micro-Raman spectrometer. The laser beam was focused on the sample surface with a spot size of 1–2 μm in confocal mode. Spectra were obtained in the 180–1,500 and 2,800–3,900 cm<sup>–1</sup> range to cover low- and high-frequency T–O stretching and vibration modes and the OH/H<sub>2</sub>O stretching regions. Spectra were obtained by measuring 3 times the same spot and 3 different spots on each sample. For each spectrum, the acquisition time was 120 s. The acquired spectra were post-processed using PeakFit<sup>®</sup> software employing a cubic baseline correction scheme similar to Di Muro et al. (2006) and Mercier et al. (2009). Quantification was obtained by acquiring Raman spectra of a total of 20 synthetic glasses with known H<sub>2</sub>O concentrations (Karl Fischer titration, KFT) ranging in composition from ol–tholeiite to rhyolite



**Fig. 1** H<sub>2</sub>O contents determined by micro-Raman spectroscopy compared with H<sub>2</sub>O contents calculated from H<sub>2</sub>O added by hydroxides or by micro-syringe. “H<sub>2</sub>O added” corresponds to the amount of H<sub>2</sub>O in the liquid phase (glass) calculated from the amount of H<sub>2</sub>O in the capsule taking into account the modal proportions of liquid and amphibole. “H<sub>2</sub>O Raman” corresponds to the H<sub>2</sub>O content determined in the recovered glasses. Analyses are reported in Table 3

and H<sub>2</sub>O contents from 0.0 to 9.8 wt%. The relative peak ratios of the OH/H<sub>2</sub>O peak(s) to either the low- (LF, 470–510 cm<sup>-1</sup>) or the high-frequency (HF, 980–1,140 cm<sup>-1</sup>) silicate vibration modes were calibrated as a function of H<sub>2</sub>O content and the LF/HF ratio (as a proxy for composition or polymerization of the silicate glass). Two calibration curves were defined, one for the range ol–tholeiite to andesite and a second for andesite to rhyolite as no single, straightforward algorithm allowed fitting over the entire compositional range. Accuracy of the method is estimated at ± 0.5 wt% H<sub>2</sub>O, and results are reported in Table 3 and displayed in Fig. 1.

## Results

Experimental conditions, phase assemblages, modes, relative Fe loss and estimated *f*O<sub>2</sub> conditions are presented in Table 2. Average phase compositions with standard deviations, number of points and Fe–Mg mineral–liquid exchange coefficients are reported in Table 4.

### General remarks

The experiments reported in this study are fractional crystallization experiments performed over restricted

temperature steps of mostly 30 °C; at lower temperatures (<830 °C), steps of 50 °C were used and two experiments (RN7W2 and RN8W2) were set between 30 °C steps, however, employing the same starting material as used for the respective lower-temperature experiments (RN7 and RN8, respectively) to explore phase relations, modes and compositions within steps that crystallized significant amounts of solids over a small temperature interval. All experiments produced large amounts of liquids (>74 %) facilitating analysis of the liquid phase using large electron beam sizes (10–20 μm) and reduced currents (7 nA) in order to minimize alkali migration in H<sub>2</sub>O-rich glasses.

A prerequisite to successfully employ the capsule-in-capsule technique to constrain the *f*O<sub>2</sub> conditions via the ferric–ferrous iron ratio in the starting material is the preservation of a near-close system in particular regarding potential iron and hydrogen loss or gain. Iron loss was evaluated either directly via the least-squares regression (LSR) analysis of the experimental charges or by a combination of LSR and image analysis to quantify the amount of Fe–Ti oxides independently (see above). The resulting iron loss was less than 6.5 % relative for all but one higher temperature experiment (RN6, 1,070 °C) that resulted 9.5 %. Fe loss of generally less than 5 % is considered insignificant close to the statistically significant limit of the detection method, and, thus, we consider the experiment as (nearly) closed system regarding iron. Hydrogen loss (or gain) by diffusion through the noble-metal capsule will significantly influence the ambient *f*O<sub>2</sub> conditions either reducing it through hydrogen gain or increasing it by loss to the outer capsule/assembly (Luth 1989). A general feature of all our experiments is the presence of vapor (fluid) bubbles (see Fig. 4) even at the highest temperature and relatively low H<sub>2</sub>O contents (3.0 wt%, RN 1), that become more frequent and generally larger with decreasing temperature and increasing H<sub>2</sub>O contents of the starting materials (up to 10.9 wt%, RN 17V2). The solubility of pure H<sub>2</sub>O considerably exceeds the amounts of H<sub>2</sub>O added to the starting material as well as the amounts calculated for the residual liquids in the respective charges. At 0.7 GPa, the solubility of H<sub>2</sub>O in basalts exceeds 8 wt% already at 0.5 GPa (at the highest run temperature of 1,170 °C) and exceeds 13 wt% at 700 °C for rhyolites for the same pressure (Newman and Lowenstern 2002; Iacono-Marziano et al. 2012), whereas the nominally added H<sub>2</sub>O contents amount to 3 wt% for the initial basalt and 12.9 wt% for the final rhyolite taking into account the liquid and hydrous phase proportions (amphibole and biotite) in lower-temperature runs. These values are considerably lower than the inferred solubility limits. The presence of bubbles is most likely caused by the presence of CO<sub>2</sub> even though this was not verified to date. We attempted to avoid any potential CO<sub>2</sub> contamination of the starting material

**Table 4** Electron microprobe analyses of experimental phases of inner/single capsules

Phases (#)	SiO <sub>2</sub>	TiO <sub>2</sub>	Al <sub>2</sub> O <sub>3</sub>	FeO <sub>tot</sub> <sup>a</sup>	MnO	MgO	CaO	Na <sub>2</sub> O	K <sub>2</sub> O	P <sub>2</sub> O <sub>5</sub>	Total	Mg# <sup>a</sup>	Fe–Mg Kd
<i>RN1, 1,170 °C</i>													
Glass (12) <sup>b</sup>	47.68 (23) <sup>c</sup>	0.69 (2)	14.27 (6)	8.32 (12)	0.18 (2)	9.64 (18)	11.50 (16)	1.85 (6)	0.61 (2)	n.m.	94.74	0.67	
<i>RN2, 1,150 °C</i>													
Glass (9)	48.15 (16)	0.69 (2)	14.71 (15)	8.33 (9)	0.18 (2)	9.12 (21)	11.77 (17)	1.78 (3) <sup>d</sup>	0.61 (3)	n.m.	95.34	0.66	
OI (14)	40.81 (18)	0.01 (1)	0.03 (3)	11.43 (12)	0.21 (2)	48.43 (24)	0.26 (3)	0.01 (1)	0.01 (1)	n.m.	101.20	0.88	0.258
Cpx (12)	53.13 (81)	0.22 (4)	3.29 (78)	4.38 (51)	0.14 (3)	17.89 (42)	21.29 (58)	0.22 (3)	0.01 (1)	n.m.	100.57	0.88	0.268
<i>RN5, 1,100 °C</i>													
Glass (30)	47.32 (16)	0.71 (3)	15.19 (50)	8.27 (23)	0.19 (3)	8.08 (39)	11.16 (17)	1.95 (7)	0.67 (4)	0.12 (3)	93.66	0.64	
OI (23)	40.43 (13)	0.01 (1)	0.03 (3)	11.70 (27)	0.26 (2)	47.50 (22)	0.26 (3)	0.00	0.00	0.01 (1)	100.20	0.88	0.241
Cpx (28)	51.45 (59)	0.32 (7)	4.20 (85)	5.53 (39)	0.16 (3)	16.37 (64)	21.51 (70)	0.27 (4)	0.02 (2)	0.01 (1)	99.84	0.84	0.330
<i>RN6, 1,070 °C</i>													
Glass (30)	47.28 (34)	0.81 (2)	17.35 (23)	7.78 (27)	0.18 (2)	6.73 (10)	10.07 (20)	2.04 (11) <sup>d</sup>	0.72 (2)	0.14 (2)	93.10	0.61	
OI (20)	39.63 (15)	0.02 (1)	0.03 (2)	16.12 (44)	0.28 (8)	44.58 (39)	0.28 (4)	0.01 (1)	0.01 (1)	0.02 (1)	100.98	0.83	0.313
Cpx (18)	51.88 (39)	0.38 (7)	3.98 (84)	5.14 (34)	0.15 (2)	16.36 (57)	21.80 (15)	0.24 (3)	0.01 (1)	0.01 (1)	99.95	0.85	0.272
<i>RN7, 1,040 °C</i>													
Glass (23)	49.99 (33)	0.95 (3)	18.74 (16)	7.53 (16)	0.17 (2)	4.68 (15)	8.15 (15)	1.59 (5) <sup>d</sup>	0.89 (4)	0.19 (3)	92.88	0.53	
Cpx (low-Al, 8)	50.68 (71)	0.46 (9)	5.33 (50)	7.96 (58)	0.25 (3)	16.50 (61)	19.15 (66)	0.28 (5)	0.00	0.02 (1)	100.63	0.79	0.300
Cpx (high-Al, 12)	46.77 (84)	0.91 (14)	10.03 (64)	8.10 (61)	0.20 (3)	13.65 (48)	20.08 (64)	0.34 (3)	0.00	0.01 (1)	100.09	0.75	0.369
Sp (11)	0.03 (3)	0.28 (4)	56.17 (64)	25.20 (86)	0.16 (2)	17.30 (54)	0.03 (2)	0.00	0.00	0.00	99.17	0.55	
Opx (5)	50.47 (60)	0.36 (2)	7.74 (80)	13.90 (16)	0.30 (5)	25.90 (10)	1.47 (13)	0.05 (2)	0.01 (1)	0.01 (1)	100.21	0.77	0.334
Plag (11)	47.65 (50)	0.03 (2)	32.05 (33)	0.56 (7)	0.01 (1)	0.09 (2)	17.14 (27)	1.81 (16)	0.06 (1)	0.02 (1)	99.42	0.22	
<i>RN7W2, 1,055 °C</i>													
Glass (20)	47.38 (22)	0.82 (3)	18.10 (13)	7.85 (14)	0.18 (3)	6.28 (8)	9.60 (15)	2.26 (9)	0.73 (3)	0.11 (3)	93.31	0.59	
Cpx (38)	49.04 (56)	0.63 (9)	7.19 (102)	7.05 (51)	0.18 (3)	14.95 (49)	20.61 (72)	0.33 (4)	0.03 (2)	0.02 (2)	100.03	0.79	0.377
<i>RN8, 1,010 °C</i>													
Glass (28)	51.62 (152)	0.67 (18)	17.89 (72)	6.45 (45)	0.17 (3)	2.65 (39)	6.62 (78)	3.00 (34)	1.14 (18)	0.21 (7)	90.42	0.42	
Opx (12)	49.81 (51)	0.37 (7)	6.44 (85)	16.17 (76)	0.37 (4)	24.66 (82)	1.66 (34)	0.03 (3)	0.02 (1)	0.02 (1)	99.55	0.73	0.269
Amph (9)	40.17 (35)	1.73 (15)	15.25 (36)	12.4 (77)	0.21 (2)	13.67 (60)	10.83 (18)	2.30 (5)	0.47 (4)	0.01 (1)	97.04	0.66	0.373
Plag (10)	46.16 (53)	0.02 (1)	32.98 (46)	0.66 (7)	0.02 (1)	0.06 (2)	17.18 (50)	1.69 (20)	0.06 (1)	0.01 (1)	98.84	0.14	
Mag (5)	0.07 (3)	4.88 (29)	14.38 (97)	71.48 (61)	0.34 (3)	4.37 (19)	0.15 (1)	0.03 (3)	0.02 (1)	0.01 (1)	95.73	0.10	
<i>RN8W2, 1,025 °C</i>													
Glass (40)	49.52 (25)	0.81 (4)	18.28 (11)	7.77 (13)	0.17 (3)	4.38 (8)	7.68 (12)	2.72 (8)	0.91 (3)	0.17 (5)	92.41	0.50	
<i>RN9, 980 °C</i>													
Glass (29)	54.81 (34)	0.57 (4)	17.57 (18)	4.28 (11)	0.17 (3)	2.60 (8)	6.67 (12)	3.00 (7) <sup>d</sup>	1.22 (3)	0.21 (5)	91.10	0.52	
Plag (20)	47.15 (86)	0.05 (4)	31.52 (105)	0.93 (24)	0.01 (1)	0.17 (15)	16.81 (76)	1.79 (28)	0.10 (6)	0.02 (2)	98.55	0.25	

Table 4 continued

Phases (#)	SiO <sub>2</sub>	TiO <sub>2</sub>	Al <sub>2</sub> O <sub>3</sub>	FeO <sub>tot</sub>	MnO	MgO	CaO	Na <sub>2</sub> O	K <sub>2</sub> O	P <sub>2</sub> O <sub>5</sub>	Total	Mg <sup>#a</sup>	Fe–Mg Kd
Mag (13)	0.03 (3)	2.23 (6)	9.00 (52)	77.18 (78)	0.47 (3)	5.56 (37)	0.11 (8)	0.00	0.01 (1)	0.01 (1)	94.60	0.11	
<i>RN10 s, 950 °C</i>													
Glass (31)	57.38 (90)	0.55 (8)	16.70 (31)	3.83 (21)	0.16 (3)	2.11 (30)	5.27 (30)	3.11 (19) <sup>d</sup>	1.41 (9)	0.25 (3)	90.77	0.50	
Amph (64)	42.61 (57)	1.96 (9)	12.95 (59)	10.18 (52)	0.26 (2)	15.44 (42)	10.88 (30)	2.14 (7)	0.40 (3)	0.03 (1)	96.85	0.73	0.363
Plag (31)	47.01 (121)	0.06 (4)	32.06 (118)	0.94 (30)	0.03 (1)	0.32 (14)	15.92 (86)	1.85 (26)	0.12 (7)	0.02 (2)	98.33	0.38	
Opx (6)	52.20 (11)	0.28 (3)	5.89 (52)	12.71 (40)	0.51 (3)	26.87 (82)	1.28 (18)	0.09 (5)	0.07 (4)	0.03 (1)	99.93	0.79	0.261
<i>RN11V2, 920 °C</i>													
Glass (30)	58.55 (34)	0.49 (3)	16.24 (12)	2.90 (8)	0.16 (3)	1.46 (6)	4.64 (10)	3.50 (7) <sup>d</sup>	1.59 (3)	0.24 (4)	89.77	0.47	
Amph (43)	42.84 (68)	2.10 (19)	12.42 (51)	10.66 (49)	0.33 (2)	15.10 (32)	10.71 (25)	2.15 (6)	0.39 (3)	0.04 (2)	96.74	0.72	0.355
Plag (19)	47.40 (181)	0.05 (3)	31.70 (119)	1.08 (41)	0.03 (2)	0.24 (13)	15.57 (98)	2.20 (50)	0.10 (5)	0.03 (2)	98.40	0.28	
Mag (15)	0.14 (11)	3.42 (15)	5.90 (62)	79.07 (128)	0.61 (3)	3.46 (26)	0.12 (4)	0.02 (2)	0.02 (2)	0.01 (1)	92.77	0.07	
Usp (3)	0.03 (3)	22.38 (24)	1.35 (3)	66.94 (25)	0.21 (2)	1.94 (7)	0.11 (4)	0.01 (1)	0.01 (1)	0.01 (1)	92.99	0.05	
<i>RN12V2–2, 890 °C</i>													
Glass (43)	61.87 (39)	0.33 (3)	15.22 (15)	1.85 (6)	0.16 (2)	1.29 (5)	3.40 (7)	3.50 (6)	1.68 (3)	0.30 (6)	89.60	0.55	
Amph (22)	43.30 (103)	1.96 (14)	11.39 (70)	10.27 (60)	0.48 (5)	15.35 (57)	10.22 (23)	1.99 (10)	0.32 (4)	0.04 (3)	95.32	0.73	0.467
Plag (15)	48.81 (244)	0.03 (2)	29.50 (188)	0.51 (16)	0.02 (1)	0.17 (10)	14.36 (200)	3.04 (106)	0.11 (4)	0.04 (3)	96.59	0.37	
Mag (21)	0.10 (9)	21.59 (218)	1.19 (11)	66.43 (211)	0.31 (5)	2.46 (30)	0.11 (3)	0.01 (1)	0.02 (1)	0.01 (1)	92.23	0.06	
<i>RN13V2, 860 °C</i>													
Glass (48)	62.30 (29)	0.25 (3)	15.29 (14)	1.64 (7)	0.15 (3)	1.06 (4)	3.33 (7)	3.22 (8)	1.83 (3)	0.23 (6)	89.30	0.54	
Amph (49)	45.97 (102)	1.19 (14)	10.86 (100)	8.94 (40)	0.59 (6)	16.86 (103)	9.96 (41)	1.78 (13)	0.32 (7)	0.05 (3)	96.52	0.77	0.343
Plag (22)	46.81 (77)	0.03 (2)	32.38 (89)	0.60 (24)	0.04 (2)	0.24 (11)	16.21 (67)	1.64 (12)	0.14 (6)	0.05 (3)	98.14	0.42	
<i>RN14V2, 830 °C</i>													
Glass (48)	63.91 (37)	0.13 (4)	14.36 (17)	0.95 (5)	0.16 (3)	0.81 (3)	2.57 (7)	3.20 (10)	1.89 (3)	0.16 (5)	88.14	0.60	
Plag (25)	46.88 (200)	0.02 (2)	31.84 (195)	0.56 (33)	0.05 (2)	0.21 (20)	15.72 (108)	1.93 (43)	0.12 (8)	0.03 (1)	97.36	0.40	
Amph (12)	46.82 (94)	1.21 (13)	9.53 (64)	9.93 (38)	0.78 (14)	16.12 (50)	9.37 (42)	1.62 (16)	0.28 (5)	0.05 (5)	95.71	0.74	0.525
Mag (18)	0.13 (12)	11.31 (227)	1.17 (13)	74.86 (221)	0.27 (7)	1.32 (29)	0.08 (2)	0.01 (1)	0.03 (1)	0.01 (1)	89.19	0.03	
<i>RN15V2, 780 °C</i>													
Glass (47)	65.90 (43)	0.09 (3)	13.29 (16)	0.68 (6)	0.12 (3)	0.51 (6)	1.60 (5)	3.73 (12)	2.11 (4)	0.10 (4)	88.13	0.57	
Amph (24)	50.51 (137)	0.73 (17)	7.36 (114)	9.23 (53)	1.25 (13)	17.34 (107)	8.19 (58)	1.20 (20)	0.23 (7)	0.03 (2)	96.07	0.77	0.399
Plag (20)	49.58 (234)	0.02 (2)	30.60 (151)	0.84 (59)	0.07 (3)	0.44 (37)	12.5 (125)	3.54 (70)	0.16 (6)	0.03 (2)	97.78	0.49	
Mag (13)	0.13 (6)	12.77 (118)	0.88 (6)	75.44 (123)	0.31 (5)	0.98 (15)	0.08 (2)	0.03 (1)	0.04 (1)	0.01 (1)	90.67	0.02	
<i>RN16V2, 730 °C</i>													
Glass (48)	67.60 (30)	0.05 (3)	11.94 (17)	0.45 (5)	0.12 (3)	0.26 (5)	0.90 (5)	3.80 (9)	2.16 (5)	0.04 (3)	87.32	0.51	
Amph (1)	54.64	0.22	2.54	9.85	3.37	20.78	4.36	0.8	0.08	0.02	96.66	0.79	0.274
Plag (13)	56.67 (152)	0.01 (1)	27.21 (112)	0.19 (15)	0.05 (2)	0.07 (7)	7.64 (114)	6.26 (53)	0.34 (9)	0.01 (1)	98.45	0.40	



**Table 4** continued

Phases (#)	SiO <sub>2</sub>	TiO <sub>2</sub>	Al <sub>2</sub> O <sub>3</sub>	FeO <sub>tot</sub> <sup>a</sup>	MnO	MgO	CaO	Na <sub>2</sub> O	K <sub>2</sub> O	P <sub>2</sub> O <sub>5</sub>	Total	Mg# <sup>b</sup>	Fe–Mg Kd
Mag (8)	0.97 (84)	12.00 (377)	0.61 (16)	74.44 (439)	0.48 (18)	0.73 (26)	0.06 (2)	0.04 (4)	0.07 (3)	0.01 (1)	89.41	0.02	
<i>RN17V2, 700 °C</i>													
Glass (50)	68.21 (37)	0.05 (2)	11.53 (15)	0.26 (3)	0.10 (3)	0.11 (2)	0.16 (3)	3.88 (9)	2.86 (5)	0.03 (3)	87.19	0.43	
Plag (6)	58.46 (142)	0.02 (1)	27.64 (199)	0.34 (21)	0.07 (5)	0.15 (7)	4.60 (66)	6.87 (31)	0.47 (7)	0.01 (1)	98.63	0.44	
Mag (2)	0.30 (4)	5.23 (153)	0.41 (5)	81.33 (158)	0.4 (11)	0.39 (1)	0.01	0.03 (3)	0.07 (2)	0.01 (1)	88.18	0.01	
Qtz (1)	99.16	0.00	0.08	0.04	0.01	0.00	0.03	0.00	0.02	0.00	99.34	0.00	
Bt (1)	38.50	1.93	16.38	7.90	0.73	18.40	0.01	0.86	8.21	0.01	92.93	0.80	0.182

<sup>a</sup> FeO<sub>tot</sub> and Mg#: all Fe as Fe<sup>2+</sup>

<sup>b</sup> Number of probe analyses

<sup>c</sup> One standard deviation of replicate analyses; 47.68 (23) should be read as 47.68 ± 0.23

<sup>d</sup> The Na<sub>2</sub>O concentration was corrected for alkali loss before normalization to 100 wt% anhydrous: RN2 I liq (1.87 wt%), RN6 I liq (2.18 wt%) and RN7 I liq (2.86 wt%), RN9 I liq (3.33 wt%), RN10 s liq (3.61 wt%) and RN11 liq (3.75 wt%)

during preparation and welding by employing wollastonite, Na<sub>2</sub>SiO<sub>3</sub> and K<sub>4</sub>Si<sub>5</sub>O<sub>12</sub> as sources for CaO, Na<sub>2</sub>O and K<sub>2</sub>O instead of commonly used carbonates and used a W-electrode instead of a graphite electrode for arc welding to prepare and shut the capsules. Figure 1 depicts the variation of the H<sub>2</sub>O contents calculated from the added amount of H<sub>2</sub>O in the starting material with the H<sub>2</sub>O content obtained by Raman spectroscopy and reveals a very consistent variation along a 1:1 correlation except for one run at 1,120 °C (RN4 not reported in Table 2, later repeated) and a possible deviation to lower but nearly constant values around 10 wt% H<sub>2</sub>O at temperatures of 860 °C and less. The high value for run RN4 (5.1 wt% Raman versus 3.04 wt% calculated) is fully consistent with the fact that this run did not crystallize any solids; therefore, it indeed contained too much H<sub>2</sub>O that was most likely introduced during preparation of the inner capsule. The upper limit of H<sub>2</sub>O contents around 10 wt% in the liquid is most probably the result of CO<sub>2</sub> contamination limiting the amount of H<sub>2</sub>O soluble in the silicate liquid. A straightforward calculation of the minimum amount of CO<sub>2</sub> that is required to saturate the experiments with a mixed H<sub>2</sub>O–CO<sub>2</sub> fluid using VolatileCalc (Newman and Lowenstern 2002, extrapolating beyond its nominal application range of 0.5 GPa) results rather consistent values of about 3,500 ppm for the basaltic liquids at highest temperatures (3 wt% H<sub>2</sub>O, 1,150 °C) and about the same amount (2,500–3,500 ppm) for rhyolitic liquids to limit the H<sub>2</sub>O solubility to 10 wt% in the liquid below 860 °C. The exsolved fluid phase is CO<sub>2</sub> rich for the basaltic to andesitic liquids (xCO<sub>2</sub> > 0.70) and becomes more hydrous (xCO<sub>2</sub> ≈ 0.5) for the felsic lower-temperature liquids. In summary, we clearly observe small amounts of vapor/fluid bubbles in all our charges that we attribute to uptake of CO<sub>2</sub> by the very fine-grained oxide–silicate starting materials. However, the amount of calculated CO<sub>2</sub> in our systems is small compared to the amount of H<sub>2</sub>O (always <5 mol % of the total volatile concentration), and we are, therefore, convinced that it does not significantly affect our phase relations, and mineral and liquid compositions. Overall the experiments were successful in maintaining the attempted H<sub>2</sub>O contents in our charges and neither gained nor lost significant amounts of hydrogen through diffusion through the noble-metal containers.

The approach to equilibrium in these experiments is facilitated by large liquid proportions always exceeding 74 %, very fine-grained starting materials (<10 μm grain size) and run durations that were optimized for equilibration versus iron and hydrogen loss that varied from 4 h at the highest temperatures to 168 h (1 week) at lowest temperatures close to the granite solidus. Elevated to high H<sub>2</sub>O concentrations further enhanced diffusional equilibration. Overheating of the starting materials to generate

homogeneous liquids was avoided to circumvent the formation of metastable phases that are reluctant to dissolve and/or re-equilibrate at lower temperatures. Approach to equilibrium can be evaluated through the following considerations: (1) homogeneity of solid and liquid phases. For the liquid phase, 9–50 individual analyses were obtained from all experiments, with >20 analyses for all but the highest temperature experiments. Liquids were analyzed throughout the entire capsule, and we did not find any statistically significant variation throughout any of the capsules. Solid phases are generally homogeneous with two exceptions: (a) plagioclase reveals a significant variation in anorthite (An) content for most of the charges (see Fig. 8) with higher An contents for the larger and well-crystallized crystals and lower An contents for small acicular to spherulitic plagioclase that most probably represent quench crystals. In addition, a few plagioclase crystals at lowest temperatures exhibit bright cores in BSE images representing residual corundum ( $\text{Al}_2\text{O}_3$ ) grains derived from the alumina oxide/hydroxide powders of the starting material (Fig. 2e) representing the only sign of incomplete reaction of starting materials observed; and (b) a single run produced sector-zoned clinopyroxene (cpx) with low-Al and high-Al sectors, a feature that is often observed in both experimental studies (e.g., Ulmer 1986; Sisson and Grove 1993a) and in field occurrences (e.g., Dessimoz et al. 2012; Hürlimann et al. 2014) in basaltic to andesitic systems and is not necessarily an expression of disequilibrium. (2) Textural observations summarized in Fig. 2 clearly reveal the presence of well-crystallized unzoned, idiomorphic solid phases in most but the lowest temperature experiments where plagioclase occasionally forms spherulitic aggregates that might either represent quench products or are related to delayed nucleation and rapid growth. (3) Major element partitioning, in particular Fe–Mg (Table 4) is very consistent with values of 0.24–0.31 for olivine, 0.27–0.33 for cpx (except the high-Al cpx, 0.37), 0.27–0.30 for opx and 0.37–0.52 for amphibole (increasing with decreasing temperature). (4) Quench crystallization is very limited to some plagioclase and rare feathery overgrowth of cpx on stable crystals that are easily identified in BSE images. Large proportions of unmodified glass were present in all experiments. For mass balance calculations, compositions of inferred quench crystals were discarded and only the average compositions of the texturally stable minerals were considered.

The attainment of near constant  $f\text{O}_2$  at the intended value corresponding to the NNO equilibrium cannot be strictly validated. However, in all experiments that contained olivine, we evaluated the  $f\text{O}_2$  utilizing a combination of  $\text{Fe}^{2+}$ –Mg partitioning assuming that olivine accommodates only ferrous iron and the known olivine–liquid  $\text{Fe}^{2+}$ –Mg partition coefficient as a function of

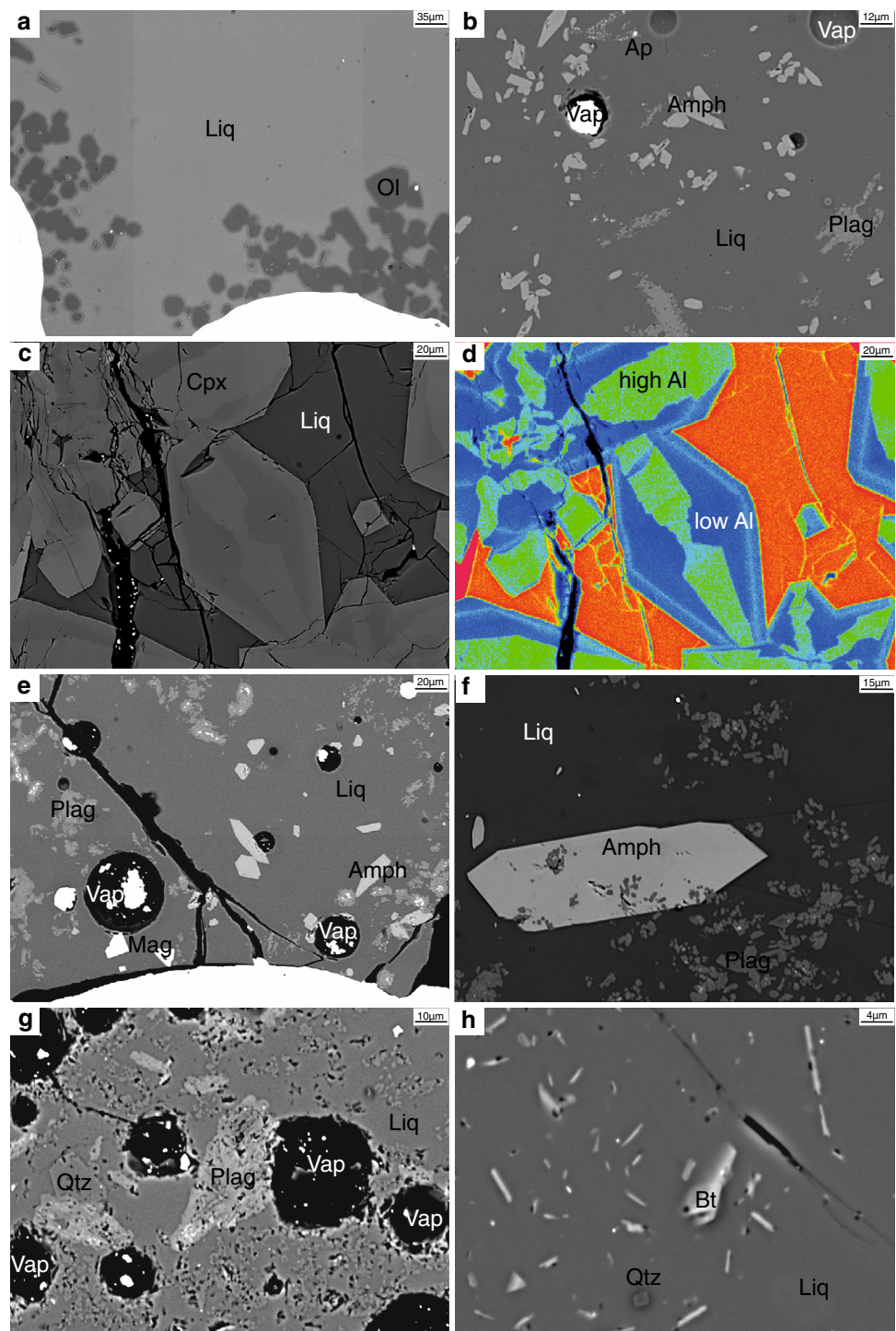
pressure using the calibration of Ulmer (1989) that has been performed on basaltic liquids using starting materials from the same dike–rock association used for the current study with the same range of  $\text{H}_2\text{O}$  contents resulting a  $\text{Fe}^{2+}$ –Mg olivine–liquid partition coefficient of 0.317. The amount of ferric iron was then estimated through a Monte Carlo type simulation taking into account the analytical uncertainty of the olivine and liquid compositions, and corresponding  $f\text{O}_2$  values were calculated using the algorithm of Kress and Carmichael (1991) relating ferric iron content of the liquid to liquid composition, temperature and pressure conditions (for detail of the method, see Kägi et al. 2005). The resulting  $f\text{O}_2$  varied between NNO and  $\text{NNO} + 1.2$  for both inner and outer capsules in the temperature range 1,170–1,055 °C (Table 2), testifying that a close approach to the targeted  $f\text{O}_2$  conditions could be achieved. As hydrogen diffusion and Fe loss presumably decrease with decreasing temperatures, we anticipate that  $f\text{O}_2$  conditions could be maintained close to or slightly above the NNO equilibrium.

#### Phase relations and phase proportions

The discussion of the phase relations and compositions is strictly limited to the results from the inner capsules in case of a double-capsule (CICT) setup as the outer capsule was generally made of alloys containing more Pd ( $\text{Au}_{80}\text{Pd}_{20}$  instead of  $\text{Au}_{90}\text{Pd}_{10}$ ) than used for the inner capsule leading to increased Fe loss in the outer capsule. In addition, the outer capsule was much larger and, thus, was subjected to a larger thermal gradient that in some cases resulted in the crystallization of additional phases at the extremities of the capsule. We provide the results obtained from the outer capsules as Electronic Supplementary Material (ESM Tables 1 and 2) as they are consistent or even equal the data obtained from the inner capsules in most cases.

Figure 3 provides a summary of the instantaneous solid compositions (ISC) rescaled to 100 % and depicts the mineral phases and their modal proportions crystallizing in every fractionation step. The accumulated percentages of the crystalline fraction (total solid fraction, TSC) from the liquidus to the actual temperature are given in addition (Fig. 4). The liquidus temperature of the initial starting composition (RC 156) was bracketed at  $1,160 \pm 5$  °C with olivine as the liquidus phase. The second phase to crystallize together with olivine is high-Ca clinopyroxene (cpx, 1,100 °C and 1,070 °C); at 1,055 °C, only clinopyroxene is present. At 1,040 °C, plagioclase and hercynitic spinel crystallize concomitantly with clinopyroxene. It is noticeable that plagioclase crystallized before amphibole. At 1,010 °C, amphibole and plagioclase with minor amounts of orthopyroxene and titanomagnetite form the solid assemblage that remains identical down temperature to

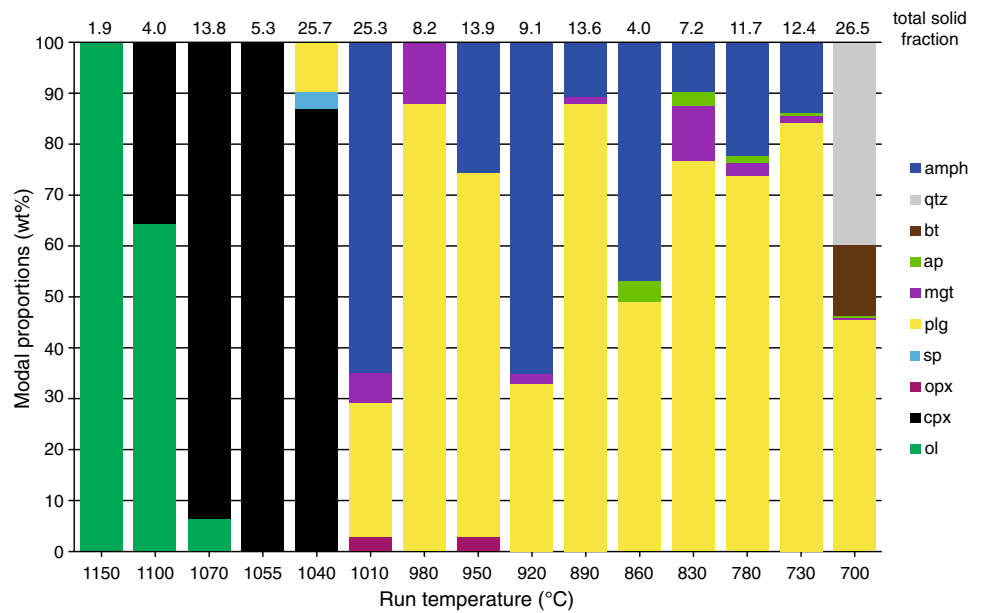
**Fig. 2** Back-scattered electron (BSE) images of representative experiments. **a** RN2—1,150 °C; **b** RN13V2—860 °C; **c** RN7—1,040 °C; sector zoning in cpx; **d** Al-K $\alpha$  X-ray distribution map of sector-zoned cpx; **e** RN12V2-2—890 °C; **f** RN10 s—950 °C; **g** RN17V2—700 °C; **h** RN17V2-2—700 °C. Abbreviations: Liq—liquid (glass), Ol—olivine, Cpx—high-Ca clinopyroxene, Plag—plagioclase, Amph—amphibole, Mag—titanomagnetite, Ap—apatite, Qtz—quartz, Bt—biotite and Vap—vapor (bubble)



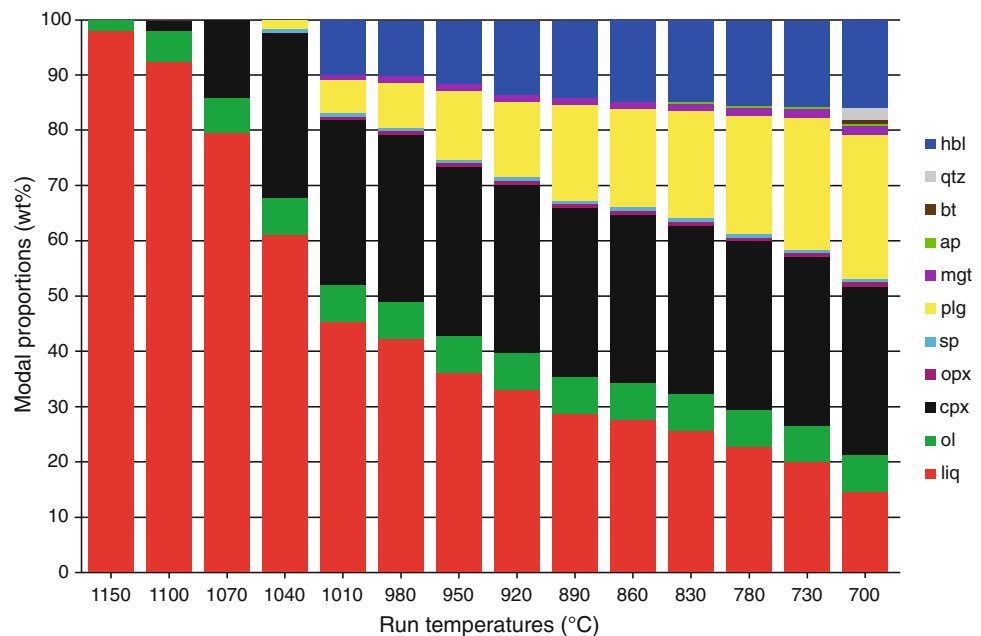
730 °C except that orthopyroxene disappears below 950 °C and apatite joins the liquidus at 860 °C. The experiment at 980 °C (RN 9) failed to crystallize amphibole and orthopyroxene that are stable at 30 °C higher and lower temperature and only contained plagioclase and magnetite as solid phases. The exact reason for this inconsistency is not clear. However, the experimental setup only approaches near-perfect fractional crystallization, a temperature step of

30 °C might leave the composition outside the stability field of some phases, but the crystallization of the stable phases, in this case plagioclase and magnetite, brings the derivative liquid ultimately back to the multiply saturated co- or peritectic curves resulting in the potential disappearance and reappearance of phases along the (incremental) experimental fractionation path. The large temperature interval from 1,010 to 730 °C produces hornblende gabbroic to

**Fig. 3** Modal proportions (wt%) of solid phases in each experiment (correspond to the instantaneous solid composition (ISC)) normalized 100 wt%. Labels on the top indicate the absolute fraction (wt%) of solids crystallized in the experiments



**Fig. 4** Total solid composition (TSC) summarizing the modal proportions of solid phases accumulated relative to the initial olivine–tholeiitic starting composition. Intermediate step experiments at 1,055 and 1,025 °C have been omitted



dioritic cumulate rocks containing minor amounts of orthopyroxene, Fe–Ti oxide (titanomagnetite to ulvöspinel) and apatite. The last fractional step produced a high-silica rhyolite liquid coexisting with plagioclase, quartz, biotite, titanomagnetite and apatite. Temperature steps with most extensive crystallization occur between 1,070 and 1,010 °C and below 730 °C. In the range 1,040–1,010 °C, two important changes in the phase relations are observed: (1) The most likely cotectic reaction where plagioclase joins the liquidus together with cpx and hercynitic spinel and (2) a peritectic reaction just above 1,010 °C where amphibole + orthopyroxene join the liquidus and replace

clinopyroxene. At 700 °C, the temperature of the granite eutectic is approached where the remaining liquid is inferred to crystallize over a small temperature interval and where K-feldspar should join the crystallizing assemblage that is composed of quartz, albite-rich plagioclase, biotite and apatite inferring another peritectic reaction where amphibole reacts with liquid to biotite terminating the crystallization interval of amphibole to lower temperature.

Figure 4 illustrates the total solid composition (TSC) representing the accumulated modal proportions of crystallized phases relative to the initial liquid (RC156). This diagram highlights the absolute amount of residual liquid

left after each fractionation step. A total of 35–40 % of ultramafic cumulates composed of dunites, wehrlites and (olivine- or plagioclase-) clinopyroxenites are extracted before gabbroic cumulate compositions occur. A total of 60 % cumulates are required to be extracted to obtain an andesitic (60 wt% SiO<sub>2</sub>) liquid at 980 °C. Only 20 % additional cumulates are formed between 980 and 730 °C, an interval of 250 °C to derive rhyolitic liquids; a somewhat unexpected result that has considerable consequences for the rheologic behavior of andesitic magmas in the Earth's crust, which is addressed in the discussion.

## Phase compositions

### Liquid

The proportions of the liquid phase present as quenched glasses were consistently high and exceeded 74 % in all experiments. For comparison and discussion, the liquid compositions were normalized to 100 % anhydrous and presented as a function of temperature in Fig. 5. The silica content is initially (1,170–1,055 °C) constant around 51 wt% and then increases abruptly at 1,040 °C followed by a continuous increase without any gaps to 79 wt% in the interval 1,025 to 700 °C. Evaluation of the extracted phase assemblages (Fig. 3) reveals that the onset of the pronounced silica increase is due to the crystallization of 3 wt% hercynitic spinel and 10 wt% anorthite-rich plagioclase. The subsequent continuous increase in silica is related to the formation of amphibole, plagioclase and minor titanomagnetite and, at higher temperature, some orthopyroxene. The final value of 79 wt% SiO<sub>2</sub> coincides with the precipitation of quartz + biotite. The saturation temperatures of clinopyroxene and plagioclase are best appreciated by inspecting the CaO and Al<sub>2</sub>O<sub>3</sub> variation diagrams (Fig. 5b and d). The bend in CaO concentration with decreasing temperature reflects the early (>1,100 °C) crystallization of clinopyroxene, and the bend in Al<sub>2</sub>O<sub>3</sub> concentration around 1,050 °C coincides with the crystallization of plagioclase. The concentrations of Na<sub>2</sub>O and K<sub>2</sub>O in the liquid (Fig. 5f and h) increase with decreasing temperature; they behave incompatible and are concentrated in the liquid throughout all the experiments. This is not perfectly true for Na<sub>2</sub>O as this oxide is included in plagioclase and amphibole, but the bulk incompatible behavior leads to an increase in the liquid with decreasing temperature. Kinks in the TiO<sub>2</sub> and P<sub>2</sub>O<sub>5</sub> concentrations (Fig. 5c and j) mark the onset of titanomagnetite and apatite crystallization at 1,010 and 860 °C, respectively. Iron oxide and magnesia concentrations (Fig. 5e and i) decrease with decreasing temperatures except for the highest temperature experiments where FeO<sub>tot</sub> concentrations remain nearly constant until a spinel phase (hercynite) joins the crystallizing assemblage. The calculated Mg-number is monotonically

decreasing until 1,010 °C and undulates toward lower temperature. We attribute this unusual behavior to the fact that Fe loss was (initially) underestimated, and therefore, the subsequent starting materials contained too low FeO<sub>tot</sub> concentration. The effect of this minor shortcoming on the phase equilibria is, however, most likely negligible as all experiments concerned are Fe–Ti oxide saturated and the total amount of FeO + MgO is rather low. MnO concentrations (Fig. 5g) remain constant until 900 °C and decrease thereafter. H<sub>2</sub>O concentrations in the residual liquids (Fig. 1) increase continuously from 3 to 10 wt% down to 860 °C and scatter around 10 wt% for the lowest temperatures either due to limited solubility caused by the presence of CO<sub>2</sub> or inability of the Raman method to accurately determine H<sub>2</sub>O contents exceeding 10 wt%.

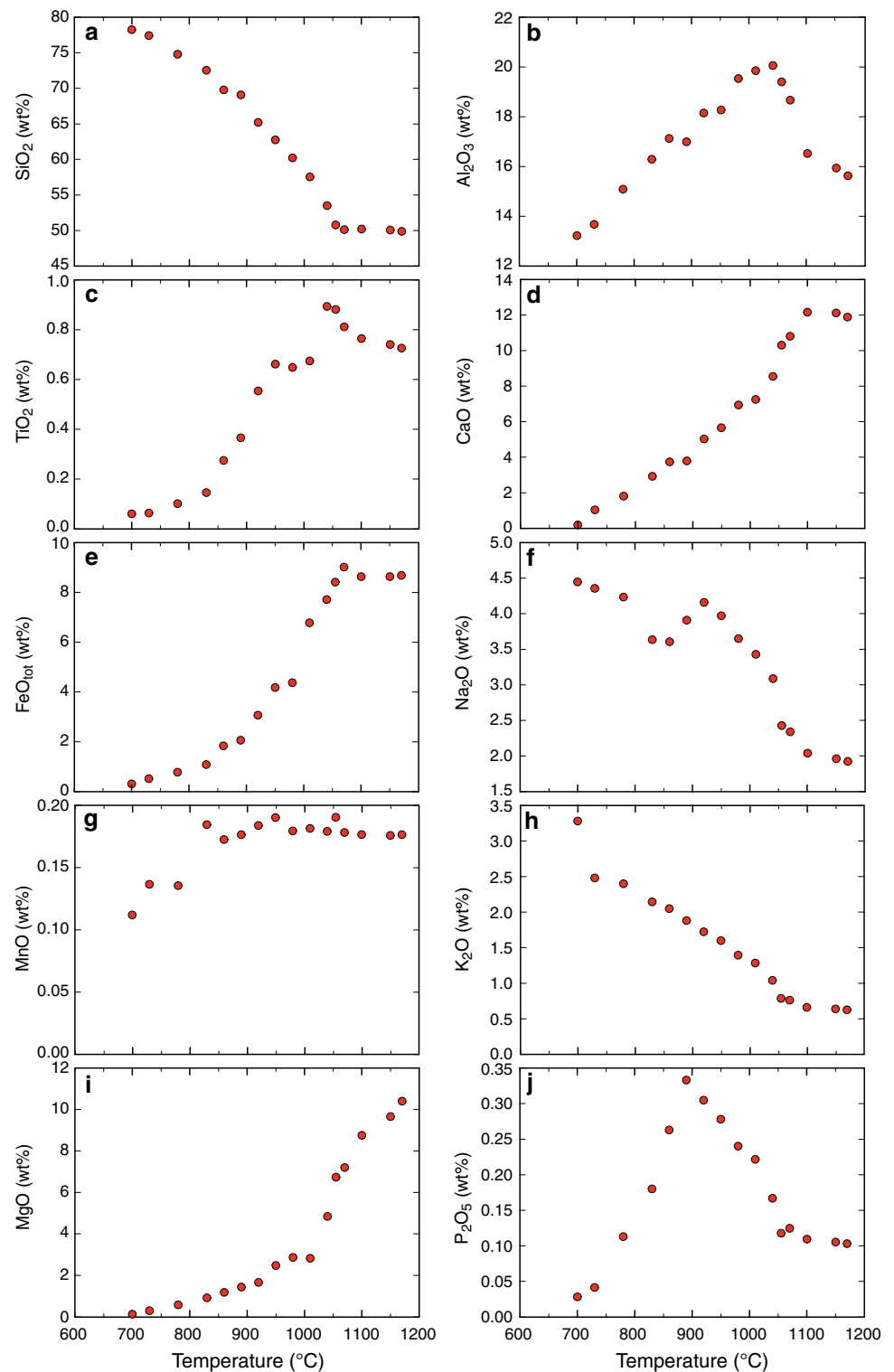
### Olivine

Olivine occurred in the range 1,150–1,070 °C, and its composition decreased from Fo 88.0 to 82.5. The Ca content of olivine is consistently low and varies between 0.24 and 0.31 wt% CaO.

### Clinopyroxene

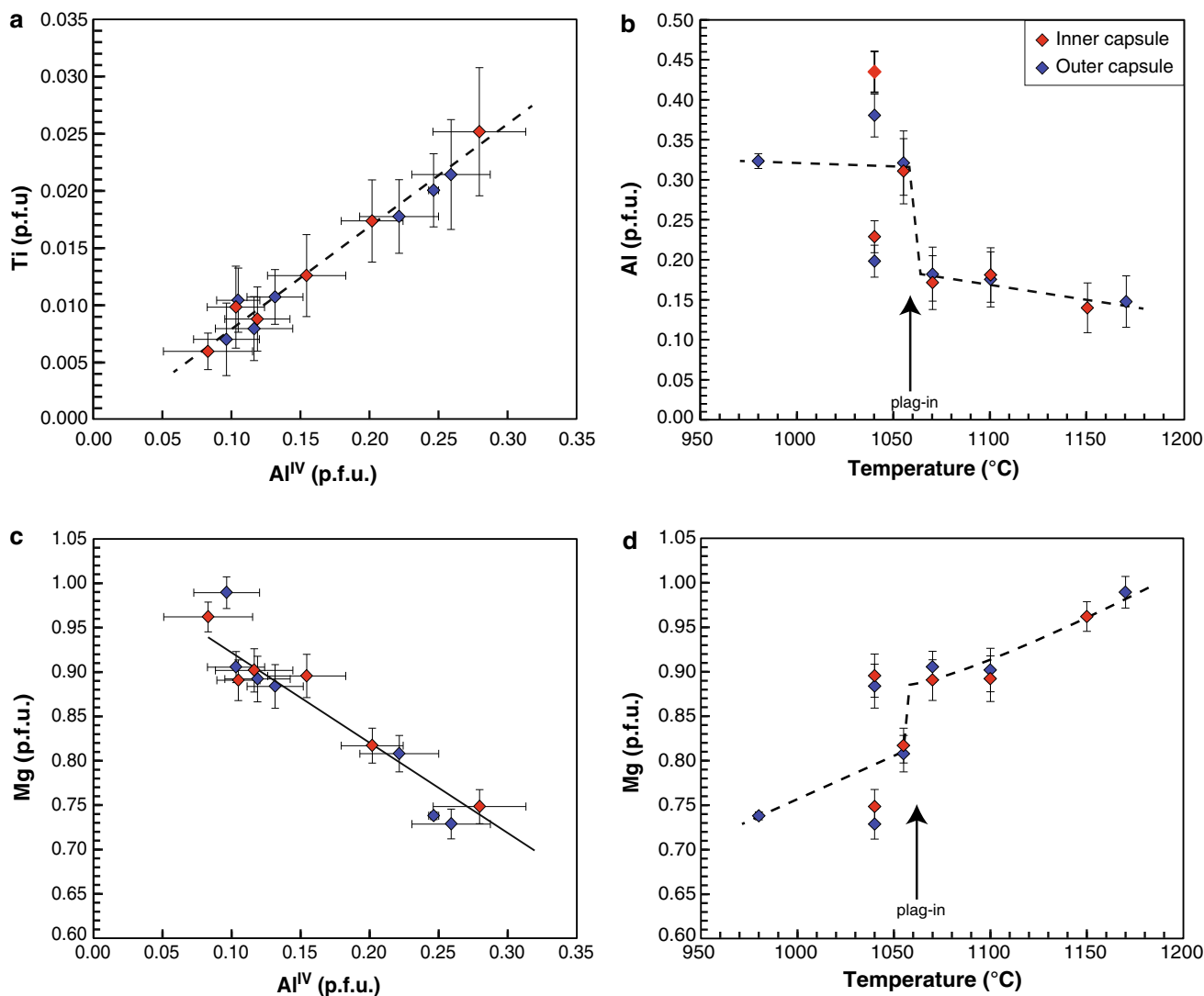
Clinopyroxene (cpx) crystallized over a wider temperature range (1,170–980 °C) in the outer capsule than in the inner capsule (1,100–1,040 °C). Crystallization of cpx additionally occurred during quench generating overgrowth rims in the form of feathery crystals on cpx that grew in equilibrium with the liquid during the experiment. The compositions of quench crystals are not considered here. All cpx from outer and inner capsules display a positive correlation of Ti versus tetrahedrally coordinated Al (Al<sup>IV</sup>, Fig. 6a) constraining Ti incorporation via the exchange  $\text{TiAl}^{\text{IV}}\text{Al}^{\text{IV}} = (\text{Mg}, \text{Fe}^{2+})\text{SiSi}$ . Similarly, Mg displays a negative correlation with Al<sup>IV</sup>, testifying that the principal exchange controlling Al incorporation in cpx is the Tschermak's vector  $((\text{Mg}, \text{Fe}^{2+})\text{Si} = \text{Al}^{\text{VI}}\text{Al}^{\text{IV}})$ ; Fig. 6c). Aluminum variation as a function of temperature (Fig. 6b) reveals several features: (1) At 1,040 °C, two distinct cpx compositions are present—one with high-Al contents (8–10 wt% Al<sub>2</sub>O<sub>3</sub>, high Ti and low Mg, Si) and one with low Al (4.5–5.5 wt% Al<sub>2</sub>O<sub>3</sub>, low Ti, high Si, Mg) denoted as low- and high-Al cpx, respectively. This is not an uncommon observation in experimental and natural systems (e.g., Villiger et al. 2004; Whitaker et al. 2008; Dessimoz et al. 2012) and represents either the expression of a stable solvus in the diopside–Tschermak's pyroxene system or, alternatively, a metastable phenomena that is not strictly limited to experimentally grown cpx in hydrous calc-alkaline systems, but likewise occurs in natural rocks ranging from plutonic (gabbroic) to hypabyssal (dike rocks) with commonly the low-Al cpx

**Fig. 5** Concentrations of major and minor oxides of the liquid (glass) phase (normalized anhydrous to 100 wt%) as a function of temperature



forming either the core of a zoned clinopyroxene (the rim being high-Al cpx) or occurring as sector-zoned cpx as shown in Fig. 2c and d (Sisson and Grove 1993a; Kawamoto 1996). (2) A conspicuous jump in Al concentration around 1,050 °C from lower to higher values that coincides with plagioclase saturation of the coexisting liquid is

evident and is most likely due to increasing Al activity in the melt phase (see Villiger et al. 2007 and references therein). The magnesium content of cpx (Fig. 6d) decreases with decreasing temperature, again changing abruptly, when plagioclase saturates, inversely correlating with increasing Tschermak's component.



**Fig. 6** Compositions of clinopyroxene as a function of  $Al^{IV}$  (a, b) and temperature (c,d). Cations per formula unit (p.f.u.) were calculated on the basis of 4 cations and 12 positive charges

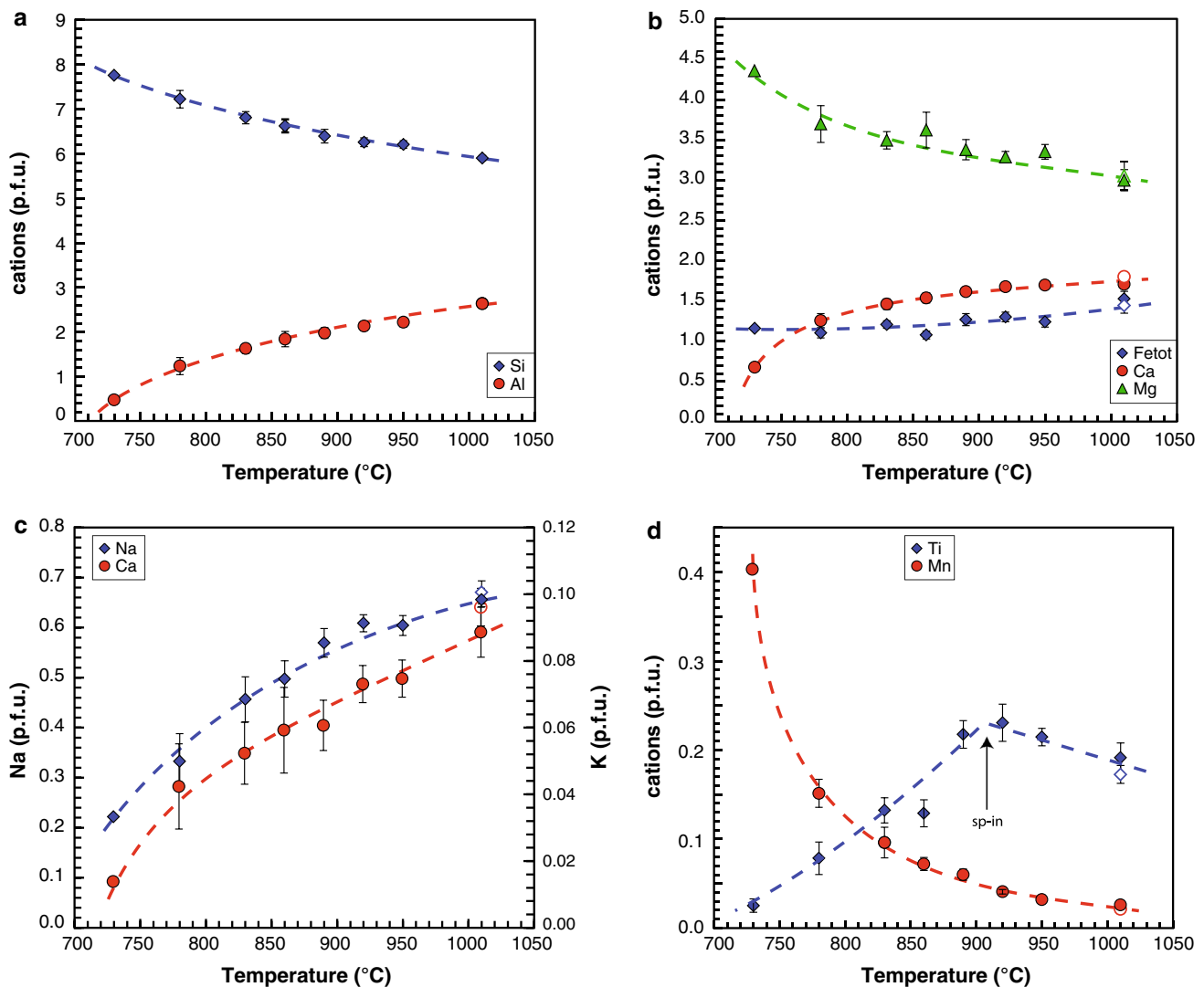
*Orthopyroxene*

The temperature range of orthopyroxene crystallization is rather narrow (1,040–950 °C). Ca, Ti and Al contents decrease with decreasing temperature; the same is true for the Mg-number except for the lowest temperature run (RN10 s, 950 °C) that showed higher Mg# for orthopyroxene consistent with the higher Mg# of the bulk system and the liquid phase.

*Amphibole*

Amphibole crystallized over a wide temperature range (1,010–730 °C). The crystals become generally smaller with decreasing temperature (Fig. 2b, e and f). Amphibole compositions have been normalized using a normalization

scheme (Spear and Kimball 1984) based on charge balance (46 positive charges without OH, F, Cl) and an empirically fixed ferric to total iron ratio ( $Fe^{3+}/Fe_{tot}$ ) of 0.33. Compositions of amphiboles span from pargasite at high temperature in basaltic–andesite liquids to magnesiohornblende in andesitic liquids to cummingtonite at the lowest temperature (730 °C) coexisting with a rhyolitic liquid. The silica content of the amphiboles monotonically increases with decreasing temperature from 5.9 to nearly 7.8 cations p.f.u. The aluminum concentration behaves opposite and decreases with decreasing temperature from 2.64 to 0.68 cations p.f.u. (Fig. 7a). The calculated tetrahedrally coordinated aluminum ( $Al^{IV}$ ) decreases strongly with temperature, whereas the octahedrally coordinated aluminum ( $Al^{VI}$ ) concentration remains nearly constant. Magnesium concentrations increase while iron concentrations remain



**Fig. 7** Composition of amphibole as a function temperature. Cations per formula (p.f.u.) unit were calculated on the basis of 46 positive charges (excluding H) and normalized with a fixed  $\text{Fe}^{3+}/\text{Fe}_{\text{tot}}$  of 0.33

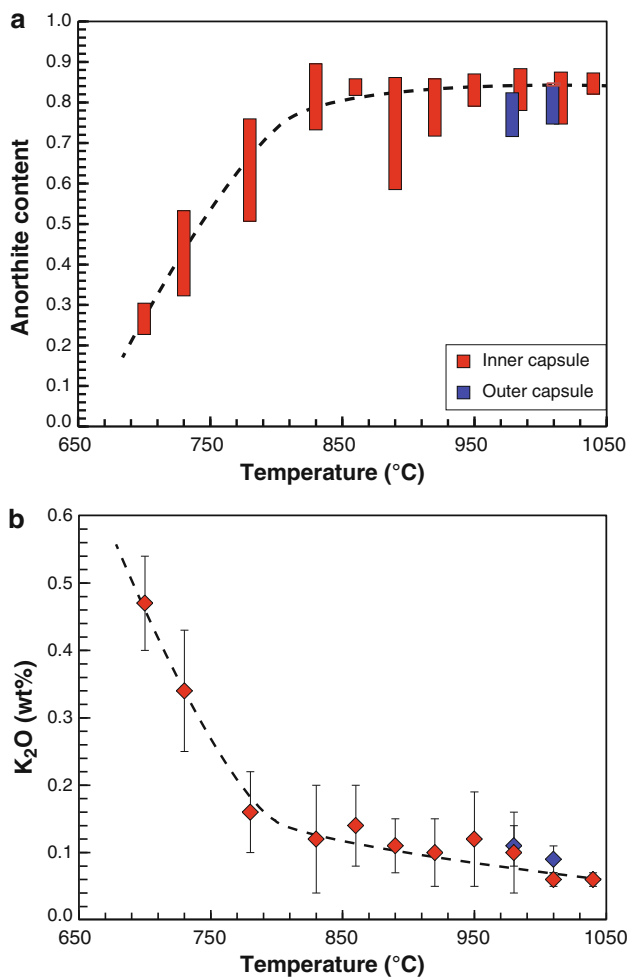
nearly constant with temperature (Fig. 7b). Calcium, Na and K concentrations (Fig. 7b and c) all decrease with decreasing temperature reflecting decreasing edenite and tschermakite and increasing cummingtonite components with decreasing temperature. Titanium slightly increases until 920 °C, and then, the trend inverts (Fig. 7d); this coincides with the onset of Ti-rich magnetite to ulvöspinel co-precipitating with amphibole strongly depleting the residual liquid in  $\text{TiO}_2$ . With decreasing temperature, Mn (Fig. 7d) is incorporated exponentially in amphibole and most probably leads to the stabilization of cummingtonite amphibole at lowest temperature.

#### Feldspar

Plagioclase crystallized from 1,040 to 700 °C. The common crystal habits are laths-shaped. Their size decreases

with decreasing temperature, and they have a tendency to form glomerocrysts (Fig. 2b, e–g). Overall, the anorthite content decreases toward lower temperatures from an average  $\text{An}_{85}$  to  $\text{An}_{27}$  but some runs exhibit rather large variations within the plagioclase population (Fig. 8a) probably due (1) early plagioclase growth around refractory corundum grains producing high An–plagioclase and (2) as evident in the microstructures, late plagioclase microlites related to quench crystallization. It is, therefore, not straightforward to determine the stable, equilibrium plagioclase composition. We prefer an overall evolution as depicted with the dashed line in Fig. 8a. Potassium concentrations remain nearly constant until 840 °C and increase thereafter to reach 0.47 wt% at 700 °C (Fig. 8b) concomitant with the strong decrease in An content and increase in  $\text{K}_2\text{O}$  of the coexisting liquid in the same temperature interval.





**Fig. 8** Anorthite (An) content (a) and K<sub>2</sub>O (wt%) concentrations (b) of plagioclase feldspar as a function of temperature. Dashed line represents the preferred evolution of the An content with temperature excluding quench microlites and high-An cores related to unreacted Al<sub>2</sub>O<sub>3</sub> oxide in the starting material

#### Iron–titanium oxides (spinel)

Various Fe–Mg–Al–Ti cubic oxides crystallized in the experiments. In the absence of any Cr in the system, the first spinel phase to saturate was hercynitic spinel with 46–56 wt% Al<sub>2</sub>O<sub>3</sub>, 25–38 wt% FeO<sub>tot</sub> and 14–17 wt% MgO at 1,040 °C, followed by Ti–Al–Mg-bearing magnetite at 1,010 and 980 °C. At 920 °C, two coexisting spinels Ti–Al–Mg-bearing magnetite (3.4 wt% TiO<sub>2</sub>) and an ulvöspinel (22.4 wt% TiO<sub>2</sub>) were observed. With decreasing temperatures, titanomagnetite becomes less TiO<sub>2</sub> rich, and finally at 700 °C, a Fe-rich Ti-bearing, Al- and Mg-poor magnetite (5.2 wt% TiO<sub>2</sub>) is stable close to the solidus. Additional information on the compositional variations of spinels among the spinel phases as a function of temperature is provided as ESM 2.

#### Apatite

Apatite crystallized in experiment at temperatures less than 890 °C as is clearly testified by the kink in the P<sub>2</sub>O<sub>5</sub> versus temperature curve (Fig. 5j). Apatite was visible in BSE images (Fig. 2b) and characteristic X-ray distribution maps (P-K $\alpha$  maps) in all runs below 890 °C, but the grains were too small to obtain reliable microprobe analysis. For estimation of the modal fractions of apatite by least-square regression, an ideal apatite composition has been used.

#### Biotite

Biotite crystallized in the lowest temperature step investigated (700 °C). The late occurrence of biotite on the liquidus might have two main reasons: (1) The compositions investigated in this study (Fig. 10) belong to a medium-K calc-alkaline series, and only at 700 °C, the K<sub>2</sub>O concentration (2.37 wt%) might be high enough to saturate a potassic phase and (2) performing near-perfect fractional crystallization experiments might delay biotite saturation by efficient removal of Fe and Mg generating rhyolitic liquids. Analyses of thin biotite laths were rather difficult as they rarely exceeded 1  $\mu$ m in thickness (Fig. 2h). Therefore, the reported biotite analyses (run RN17) should be considered indicative rather than precise.

#### Quartz

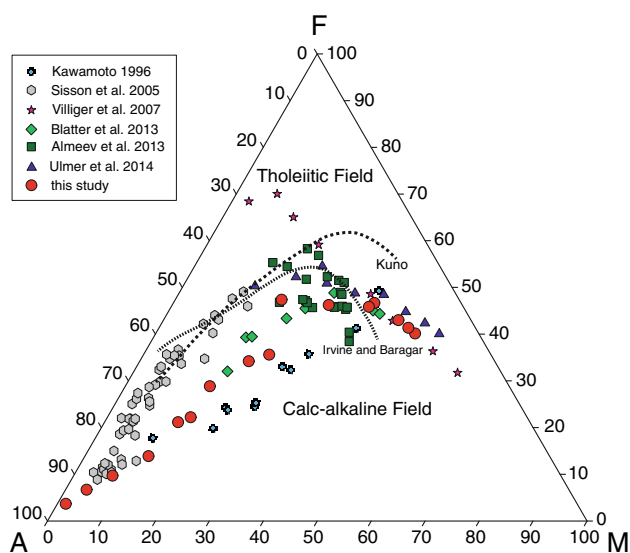
Quartz crystallized in the last fractionation step together with biotite, plagioclase, magnetite and apatite when the SiO<sub>2</sub> concentration in the liquid reached 79 wt%. Crystals are idiomorphic and most of them less than 1  $\mu$ m in size (Fig. 2g). Due to their idiomorphic shape, it can be excluded that they represent non-reacted SiO<sub>2</sub> from the starting material.

## Discussion

#### Comparison of liquid compositions with previous studies

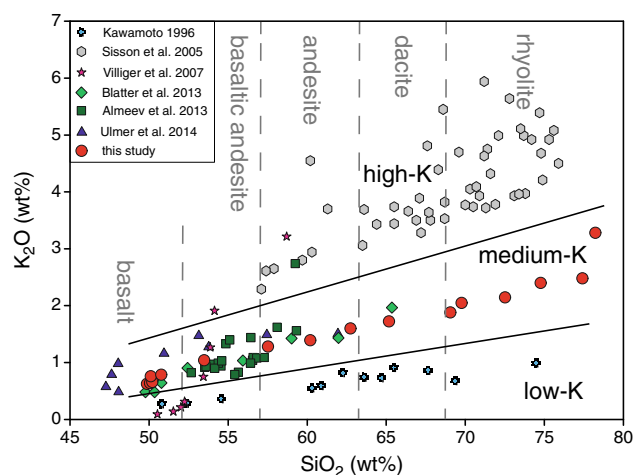
The first part of the discussion presents a comparison of the experimentally derived liquid compositions of this study with those from previous studies on hydrous calc-alkaline and anhydrous (tholeiitic) differentiation at a pressure of 0.7 GPa.

In order to distinguish the tholeiitic from the calc-alkaline differentiation series, the AFM diagram of Wager and Deer (1939) (Fig. 9) with discrimination lines according to Kuno (1968) and Irvine and Baragar (1971) was utilized. The liquid compositions of this study clearly



**Fig. 9** AFM (A = Na<sub>2</sub>O + K<sub>2</sub>O, F = FeO, M = MgO; wt%) diagram discriminating tholeiitic and calc-alkaline magmatic series. Separation lines according to Kuno (1968) and Irvine and Barager (1971). Data sources and experimental conditions: Kawamoto (1996)—high-alumina basalt, 2 wt% H<sub>2</sub>O, equilibrium crystallization, NNO +1.3, 0.5 GPa; Sisson et al. (2005)—basalt, 1.7–2.3 wt% H<sub>2</sub>O, equilibrium crystallization, NNO –1.3 to +4, 0.7 GPa; Villiger et al. (2007)—ol–tholeiite, anhydrous, fractional crystallization, C–CO<sub>2</sub>, 0.7 GPa; Blatter et al. (2013)—Mg–basalt, 2.1 wt% H<sub>2</sub>O, equilibrium crystallization, Re–ReO<sub>2</sub>, 0.7 GPa; Almeev et al., (2013)—basaltic andesite, 3.8–10.7 wt% H<sub>2</sub>O, equilibrium crystallization, FMQ +2.7 – FMQ +4.1, 0.7 GPa; Ulmer et al. (2014)—picobasalt, 2.6 wt% H<sub>2</sub>O, fractional crystallization, NNO +0.4, 1.0 GPa

follow a calc-alkaline trend. This is distinctly different from the study of Villiger et al. (2007) showing enrichment in iron and, therefore, follows a tholeiitic trend. It should be noted that both studies have a comparable initial liquid composition. Villiger et al. (2007) conducted anhydrous fractionation experiments under more reducing conditions (graphite-saturated), and this study simulates hydrous fractional crystallization at  $f_{O_2}$  conditions close to the NNO equilibrium. Our results are consistent with other studies in hydrous systems (Kawamoto 1996; Sisson et al. 2005; Almeev et al. 2013; Blatter et al. 2013) that follow a calc-alkaline trend. The starting compositions of those studies are somewhat different than the one employed in this study ranging from basaltic to basaltic andesitic (Sisson et al. 2005) to high-alumina basalt (Kawamoto 1996). Sisson and Grove (1993b) showed that high-alumina basalt can be reached by fractional crystallization from hydrous basalt. Almeev et al. (2013) used a basaltic andesite with variable H<sub>2</sub>O–CO<sub>2</sub> (xH<sub>2</sub>O from 0.0 to 1.0) ratios, which was generally fluid saturated. Blatter et al. (2013) investigated a Mg–basalt similar to the initial starting material (RC156c) used in our experiments, but with higher TiO<sub>2</sub>, FeO<sub>tot</sub> and Na<sub>2</sub>O concentrations and lower Al<sub>2</sub>O<sub>3</sub> and MgO contents and a lower H<sub>2</sub>O content of 2.1 wt%. The



**Fig. 10** K<sub>2</sub>O–SiO<sub>2</sub> (wt%) diagram to discriminate low-, medium- and high-K calc-alkaline series. Diagram modified after Gill (1981) by extension to higher SiO<sub>2</sub> concentrations. Compositional boundaries of rock names used in this study for liquid compositions are indicated by gray labels and dashed lines. For data sources and experimental conditions, see Fig. 9

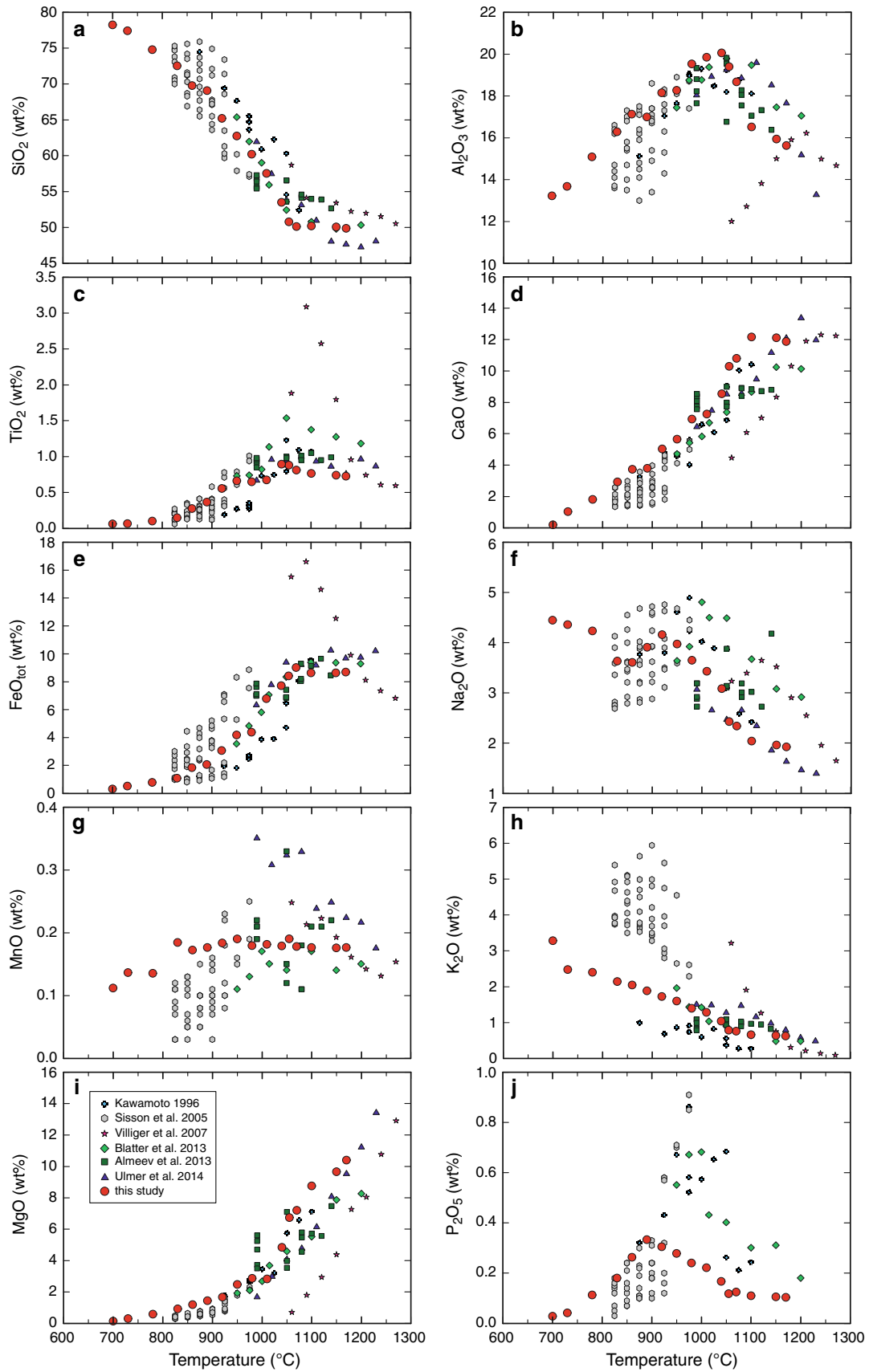
difference in starting composition seems not crucial with respect to the calc-alkaline trend as long as they contain H<sub>2</sub>O. In Fig. 10, the K<sub>2</sub>O–SiO<sub>2</sub> diagram after Gill (1981) used to subdivide the calc-alkaline series further into low-, medium- and high-K series, the various experimental studies plot at rather different locations in this diagram, identifying the series of Kawamoto (1996) as low K, Sisson et al. (2005) as high K and this study, Almeev et al. (2013) and Blatter et al. (2013) as medium K. The different evolutionary trends are clearly related to the respective character of the starting material employed in the different studies, emphasizing the importance of the choice of the starting material as well as evidently putting some limits on the application of the experimental data to natural systems.

Figure 11 compares the liquid compositions obtained by the various studies. It should, however, be noted that all previous studies on hydrous systems at comparable pressures were crystallization experiments on fixed bulk compositions, i.e., equilibrium crystallization experiments and, thus, are not strictly comparable with the present experimental data set that additionally covers a much wider temperature and/or composition interval from a near-primary, hydrous olivine–tholeiite to high-silica rhyolite along a single liquid line of descent (LLD). For comparison, the studies of Kawamoto (1996), Sisson et al. (2005), Villiger et al. (2007), Almeev et al. (2013) and Blatter et al. (2013) were selected to highlight compositional differences of (1) tholeiitic to calc-alkaline series and (2) the importance of H<sub>2</sub>O and  $f_{O_2}$ . Additionally, we utilize data from a similar fractional crystallization study conducted at 1.0 GPa starting from a hydrous picobasalt (2.6 wt% H<sub>2</sub>O, 47 wt% SiO<sub>2</sub>, 17 wt% MgO) originating

from the same dike suite from the Tertiary Adamello batholith (Ulmer et al. 2014) to evaluate the pressure effect on the liquid line of descent. The liquidus temperature of the anhydrous tholeiitic system (Villiger et al. 2007) is higher for comparable starting compositions, a well-known effect and clearly related to the presence of H<sub>2</sub>O in the hydrous system. Major compositional differences are the silica content that only moderately increases in the anhydrous tholeiitic series compared to the calc-alkaline series. The rather abrupt increase in SiO<sub>2</sub> concentration coincides with the onset of spinel (hercynite and/or Ti-magnetite) and An-rich plagioclase crystallization (this study). The SiO<sub>2</sub> enrichment is then progressively continued with the co-precipitation of titanomagnetite, plagioclase and amphibole. Similar behavior is reported by the data presented by Kawamoto (1996), Almeev et al. (2013) and Blatter et al. (2013) and is intimately linked in the latter cases to extensive amphibole crystallization. The 1.0-GPa fractional crystallization experiments (Ulmer et al. 2014) reveal a similar behavior as the 0.7-GPa fractional data with abruptly increasing silica contents at 1,140 °C linked to the crystallization of abundant (5 wt%) hercynitic spinel over a small temperature interval. Considering all data sets together, the silica increase typical for calc-alkaline differentiation is not caused by a single phase occurring on the liquidus, but depending on bulk composition, fO<sub>2</sub> and fH<sub>2</sub>O and pressure conditions, it is caused either by the onset of significant Fe–Ti–Al oxide crystallization with or without An-rich plagioclase or amphibole. Common to all these scenarios is the suppression of extensive plagioclase crystallization occurring in anhydrous systems and a shift of the plagioclase compositions for a given Ca/Na ratio of the system to higher An-contents as pointed out and quantified by Sisson and Grove (1993a). Alumina (Fig. 11b) reaches considerably higher concentrations in hydrous experiments due to the suppression of plagioclase saturation toward lower temperature. This effect of H<sub>2</sub>O attributed to its depolymerizing effect on the liquid structure and thus suppression of highly polymerized phases at the expense of less polymerized phases (olivine and pyroxene) is well known and has been addressed by a number of studies (e.g., Beard and Lofgren 1991; Gaetani et al. 1993; Kawamoto 1996). The various studies conducted in hydrous systems all show significant increase in Al<sub>2</sub>O<sub>3</sub> reaching 18–20 wt%, the present one and the comparable study at 1.0 GPa reaching the highest concentrations despite the fact that the starting materials contain the lowest Al<sub>2</sub>O<sub>3</sub> concentrations (Ulmer et al. 2014). In all studies considered here, clinopyroxene occurs on the liquidus shortly after olivine or is actually the liquidus phase (Kawamoto 1996; Sisson et al. 2005; Almeev et al. 2013) as clearly evident from the CaO evolution (Fig. 11d) that monotonically decreases with the exception

of Blatter et al. (2013) and Ulmer et al. (2014) that show a small interval of CaO increase where olivine (± Cr-rich spinel) is the only liquidus phase. The behavior of TiO<sub>2</sub> (Fig. 11c) is controlled by Fe–Ti oxides (and at lower temperature also amphibole) and is dependent on initial TiO<sub>2</sub> concentration, H<sub>2</sub>O content and fO<sub>2</sub>. Villiger et al. (2007) saturate ilmenite (fO<sub>2</sub> buffered by C–CO<sub>2</sub>, below FMQ at 0.7 GPa, Ulmer and Luth 1991), which decreases not only the titanium but also the iron concentration in the liquid significantly. In more oxidized systems (FMQ to Re–ReO<sub>2</sub>), the prevalent Ti-bearing phase is magnetite–ulvöspinel solid solution. Amphibole incorporates also a significant amount of titanium, and its modal abundance is higher than that of Fe–Ti oxides. The FeO<sub>tot</sub> concentration in hydrous systems is monotonically decreasing unlike in the tholeiitic system (Villiger et al. 2007) due to increased H<sub>2</sub>O activity, which depresses the stability of silicate minerals relative to Fe–Ti oxides (Beard and Lofgren 1991; Sisson and Grove 1993a). The overall trends for most oxide concentrations for hydrous experiments around 0.7 GPa are consistent with each other; some differences (e.g., K<sub>2</sub>O and potentially also P<sub>2</sub>O<sub>5</sub>) are clearly related to differences in the starting composition preferentially affecting the most incompatible elements. The fundamental difference between the fractional crystallization experiments conducted in this study compared with equilibrium crystallization experiments performed in previous studies is that in our study, the evolution of the intermediate and acidic differentiates is not “preconditioned” by the choice of the starting material except for the most incompatible elements. Another important difference is that peritectic reactions (see below) in particular olivine + liquid = opx and the amphibole-forming reactions involving pyroxenes ± plagioclase ± Fe–Ti oxide are suppressed in the present study, but in equilibrium experiments often limit the compositional range of differentiates that can be attained for a given starting composition.

A comparison of the results of the experimental study with thermodynamic calculations using MELTS (Ghiorso and Sack 1995; Asimow and Ghiorso 1998) is provided as Electronic Supplementary Material (ESM 3) revealing that fractional crystallization calculations employing the Rhyolite-MELTS code (Gualda et al. 2012) fails to predict some of the typical characteristics such as strong silica enrichment and results alumina contents far in excess of the observed ones along the liquid line of descent. This is not surprising as it is clearly stated on the MELTS homepage (<http://melts.ofm-research.org/> December 2013) that “phase equilibria involving hornblende and biotite are not modeled well by the MELTS package and consequently simulating the evolution of intermediate to silica-rich calc-alkaline systems is not recommended.”



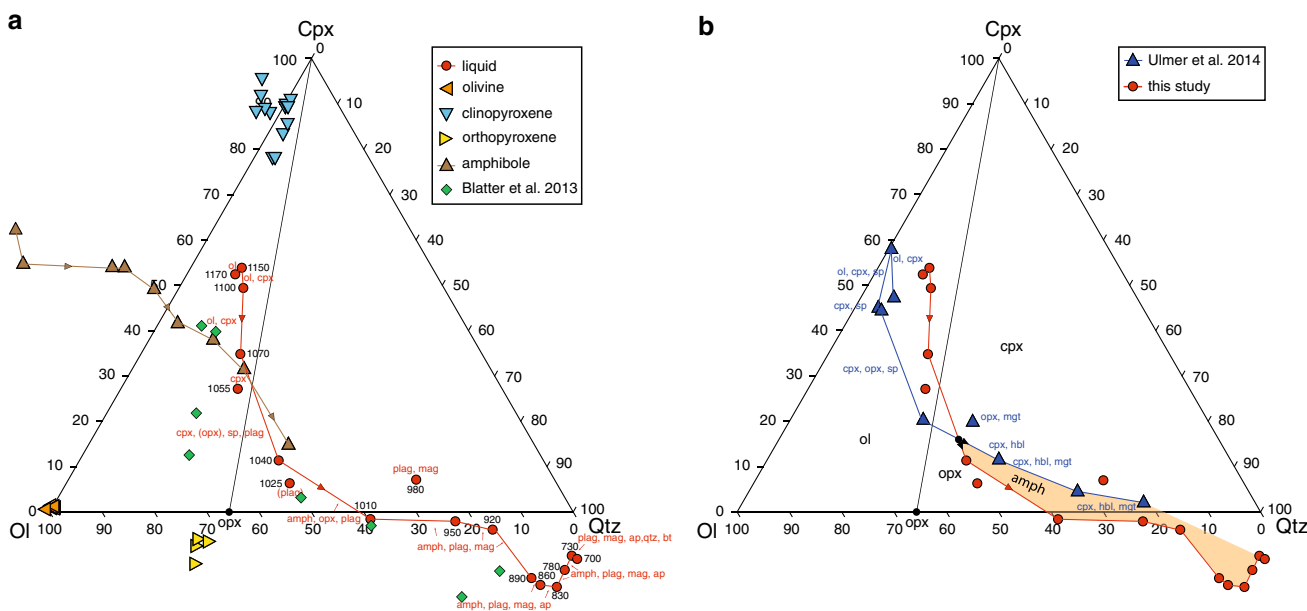
**Fig. 11** Comparison of liquid oxide concentrations (wt%) of this study with other experimental studies at similar conditions as a function of experimental temperature. For data sources and conditions, see Fig. 9

Liquid lines of descent and phase relations in ternary projections

Figure 12a provides the compositional variations of the liquids and selected mineral phases of the 0.7 GPa experiments in a pseudoternary Ol – Cpx – Qtz diagram projected from plagioclase, orthoclase and Fe–Ti–Cr oxides after Grove et al. (1992). Temperatures (Fig. 12a) and coexisting solid phase assemblages are specified along the liquid lines of descent for each melt projection point. The highest temperature experiments of this study are dominated by crystallization of olivine and subsequent clinopyroxene. The liquid line of descent (LLD) moves toward the Cpx corner as long as olivine is the dominating phase, but turns sharply when the proportions of clinopyroxene are becoming dominant moving away from the cpx–olivine sideline toward the Ol–Qtz sideline. The LLD crosses the Ol–Qtz sideline at about 1,010 °C becoming corundum-normative (peraluminous, negative Cpx coordinate) coinciding with the onset of amphibole fractionation and concomitantly with plagioclase (projecting phase) and small amounts of opx and titanomagnetite. At this point, the LLD bends toward the Qtz corner principally driven by

amphibole fractionation. The amphibole compositions are provided for reference; they change their compositions from nepheline-normative (negative Qtz coordinate) to Qtz saturated (inside the Cpx–Ol–opx subtriangle) to finally Qtz oversaturated (Cpx–Qtz–opx) with decreasing temperature. From 1,010 to 730 °C the LLD follows and outlines the amphibole stability field. The final bend at low temperatures toward the Qtz corner is most likely a computational artifact caused by the large amount of Qtz in the calculated normative mineral modes with small amounts of negative Cpx and Ol (with large amounts of plag not considered as it is a projection phase); the liquids, however, strictly remain peraluminous.

Figure 12b compares the LLD of the present study with that of Ulmer et al. (2014) at 1.0 GPa who started the fractionation experiments with a picrobasaltic composition containing a considerably higher normative olivine content, and therefore, their liquid line of descent starts closer to the Ol–apex compared to the olivine–tholeiitic starting composition employed in this study. The two LLDs at 0.7 and 1.0 GPa converge on their way toward the Ol – Qtz sideline. This convergence is enhanced because in the 1.0 GPa experiments olivine and cpx crystallize more abundantly (generating nearly 50 % dunitic, wehrlitic and (clino-)pyroxenitic cumulates) deflecting the LLD more effectively toward the Qtz corner. The crossing point (marked with a black dot) of both LLDs coincides with the onset of amphibole crystallization: at 0.7 GPa, amphibole and



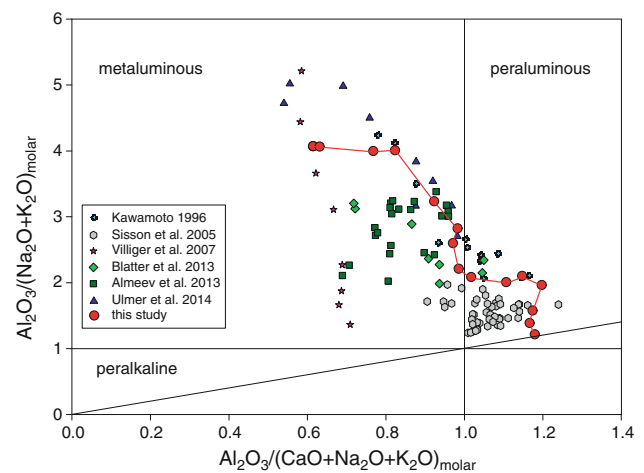
**Fig. 12** Normative molecular, pseudoternary olivine (Ol)–clinopyroxene (Cpx)–quartz (Qtz) projections using the projection scheme of Grove et al. (1992) projected from plagioclase, orthoclase and Fe–Ti–Cr oxide for **a** liquid (glass), selected solid phases of this study and liquid compositions from Blatter et al. (2013); **b** comparison of liquid

compositions from this study with liquids obtained in the fractional crystallization experiments of Ulmer et al. (2014) conducted at 1.0 GPa. Labels in a denote run temperatures and in b coexisting solid assemblages

orthopyroxene crystallize concomitantly with plagioclase and Ti-magnetite, whereas at 1.0 GPa, amphibole coexists with clinopyroxene and magnetite. The crossover of the two LLDs at 0.7 and 1.0 GPa is evidently related to (1) the coexisting parageneses: amphibole + orthopyroxene + plagioclase (the projection phase of Cawthorn and O'Hara 1976) at 0.7 GPa versus amphibole + clinopyroxene at 1.0 GPa; and (2) the likely existence of a diverging peritectic point encompassing the amphibole stability field with the orthopyroxene–amphibole (+plagioclase +Ti-magnetite) forming the lower branch and clinopyroxene–amphibole (+Ti-magnetite) forming the upper branch (Fig. 12b). A similar reaction scheme, termed “distributary reaction,” was previously proposed for the appearance of amphibole at 0.2 GPa involving olivine and clinopyroxene instead of clinopyroxene and orthopyroxene (Sisson and Grove 1993a, b). The exact path at amphibole saturation potentially depends on several parameters: (i) pressure (and Na<sub>2</sub>O and H<sub>2</sub>O content of the initial liquid) suppressing plagioclase stabilization at 1.0 GPa potentially leading to more calcic liquids saturating clinopyroxene instead of orthopyroxene, (ii) silica activity that is probably higher in the 0.7 GPa experiment due to less orthopyroxene being extracted in the steps prior to amphibole saturation, favoring orthopyroxene crystallization at lower temperatures and (iii) generally more extensive crystallization at 1.0 GPa (lower melt fraction left) before the onset of amphibole crystallization than at 0.7 GPa. Theoretically, peritectic reactions should only be recorded by a change in the crystallizing mineral assemblage in perfect fractional crystallization in a stepwise approach, always oversaturating the liquid coexisting with solids at a 30 °C lower temperature. Therefore, peritectic reactions will definitely influence the crystallization paths (LLD) during crystallization-driven differentiation.

The generation of peraluminous differentiates

The contrasting phase assemblages observed at 0.7 and 1.0 GPa at the onset of amphibole crystallization leads to the evolution of the liquids at 0.7 GPa into the corundum-normative peraluminous field (negative cpx-component) at a SiO<sub>2</sub> content of 62 wt%. The LLD at 1.0 GPa ends at 990 °C with a SiO<sub>2</sub> content of 62 wt% (volatile-free basis) just inside the metaluminous field, its continuation most likely will also cross the metaluminous–peraluminous join assuming amphibole crystallization is continued. The LLD of this study has been extended to 700 °C to explore the compositions of differentiated dacitic to rhyolitic liquids that are all peraluminous (Fig. 13) and, except for the last fractionation step at 700 °C, coexist with amphibole, relatively Ca-rich plagioclase and titanomagnetite. This is the



**Fig. 13** Alumina saturation (ASI) versus peralkalinity (PI) index diagram modified after Maniar and Piccoli (1989) discriminating metaluminous, peraluminous and peralkaline liquids. For data sources, see Fig. 9

most typical gabbroic assemblage found in intermediate to deep-crustal hydrous, calc-alkaline plutonic rocks (e.g., Davidson et al. 2007; Larocque and Canil 2010; Dessimoz et al. 2012) and testifies that differentiation from intermediate (andesitic) to acidic calc-alkaline magmas involves extensive crystallization of amphibole–gabbros.

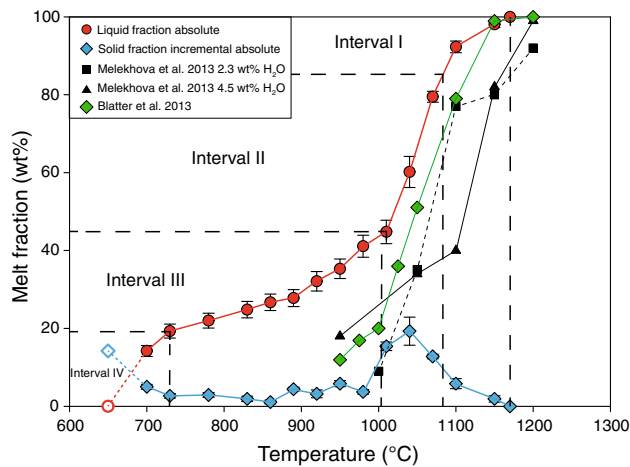
The aluminum is concentrated in the hydrous liquid as a consequence of the enhanced clinopyroxene and amphibole crystallization at elevated pressure (0.7 GPa) and H<sub>2</sub>O contents relative to orthopyroxene and plagioclase under lower pressure and/or lower H<sub>2</sub>O contents. Liquid compositions reach peraluminous compositions (Fig. 13) at SiO<sub>2</sub> contents of 65.2 wt% (62.3 wt% (Kawamoto 1996); 64.4 wt% (Sisson et al. 2005) and 60.5 wt% (Blatter et al. 2013)). Cawthorn et al. (1976) already stated that the exact value of the SiO<sub>2</sub> content at which the liquids become corundum-normative varies considerably. Sisson et al. (2005) reported only metaluminous liquid compositions for the most reducing experiments. For comparison, the anhydrous tholeiitic series of Villiger et al. (2007) is strictly metaluminous, and there is no indication that even more differentiated liquids would approach the metaluminous–peraluminous boundary; the principle difference being the more prominent role of orthopyroxene in the fractionating assemblage. As proposed by several studies, experimentally and field based (e.g., Cawthorn et al. 1976; Müntener et al. 2001; Davidson et al. 2007), the formation of corundum-normative intermediate to acidic (dacite to rhyolite) magmas does not a priori require contamination by or assimilation of partial melts derived from pelitic crustal lithologies as anticipated in numerous studies (Wall et al. 1987; Clemens 1989; Clemens et al. 2011), but can be achieved by fractional crystallization of hydrous basaltic

parental magmas in mid to lower crustal (>0.6 GPa) magma reservoirs. At least in the case of the Adamello batholith, corundum-normative tonalites and granodiorites systematically show REE patterns that are indicative for extensive amphibole and very late plagioclase fractionation (e.g., Ulmer 2007). Interestingly, these are the least crustally contaminated rocks in the entire batholiths with the lowest  $\epsilon\text{Sr}$  and highest  $\epsilon\text{Nd}$  (Kagami et al. 1991; Hürliemann et al. 2014) and  $\epsilon\text{Hf}$  (Broderick 2013) values, approaching pristine mantle values, thus supporting the derivation of corundum-normative intermediate to acidic magmas by crystallization dominated processes at mid to lower crustal conditions instead of invoking large amounts of crustal contamination.

#### Melt fraction as a function of temperature and some rheologic consequences

Figure 14 displays a diagram illustrating melt fractions relative to the initial mass (ol–tholeiite RC156c) and the solid fraction extracted in each temperature step relative to the initial magma mass as a function of temperature. The observed relationship is distinctly nonlinear: A first interval between the liquidus around 1,170 °C and 1,100 °C (I) is characterized by minor (<10 %) olivine and clinopyroxene extraction, between 1,070 and 1,010 °C 50 % of the original mass crystallizes to clinopyroxene with minor plagioclase and hercynitic spinel followed by amphibole plus plagioclase plus minor opx. The peritectic reaction at 1,020 °C does not result in a significant increase in the solid fraction, because fractional crystallization suppresses the peritectic reaction that is just reflected by a change in the crystallizing assemblage from clinopyroxene-dominated to amphibole + plagioclase-dominated solid assemblages. This massive crystallization interval (II) prior to and at the onset of amphibole crystallization is also responsible for effectively driving the liquid composition from basalt to andesite. The following large temperature interval (III) of 280 °C toward the granite eutectic below 700 °C (not determined in this study, but located by Stern et al. (1975) around 650 °C for  $\text{H}_2\text{O}$ -saturated conditions) is characterized by much reduced crystallization of a single paragenesis composed of amphibole–plagioclase–magnetite with accessory apatite joining around 860 °C. The solid fraction extracted over this entire 280 °C interval only amounts to about 20 % (relative to the starting olivine–tholeiitic liquid), leaving about 20 % of the original mass as high-silica rhyolite that solidifies at the granite–eutectic minimum over a small temperature interval (650–700 °C; Stern et al. 1975). To emphasize the importance of fractional versus equilibrium crystallization, Fig. 14 shows in addition the results of equilibrium crystallization experiments at an identical pressure (0.7 GPa) on similar near-primary,

arc-related, hydrous basalts from Blatter et al. (2013) and Melekhova et al. (2013) for two initial  $\text{H}_2\text{O}$  contents (2.3 and 4.5 wt%) that bracket the 3.0 wt% of the initial, basaltic starting material of this study. The evolutionary trends of the equilibrium crystallization experiments are significantly different from the trends obtained in this study: The melt fraction and the evolution to more differentiated (silica rich) compositions span a much smaller temperature interval: Melekhova et al. (2013) with 2.3 wt%  $\text{H}_2\text{O}$  already obtained at 1,000 °C a granitic liquid (75.1 wt%  $\text{SiO}_2$ ) at a melt fraction less than 10 %, whereas in the present study the liquid at 1,010 °C is andesitic (56.6 wt%  $\text{SiO}_2$ ) at a melt fraction of 45 %. With an initial  $\text{H}_2\text{O}$  content of 4.5 wt%, a melt fraction of 35 wt% and an andesitic liquid (56.9 wt%) exists in the equilibrium crystallization experiments, whereas in the fractional case, at the same temperature, 75 % liquid of basaltic composition (50.8 wt% silica) is present. Likewise, Blatter et al. (2013) obtain a dacitic liquid (62 wt%  $\text{SiO}_2$ ) representing 17 wt% melt fraction at 975 °C. It is obvious that the suppression of peritectic reactions results in both much smoother melt fraction–temperature as well as melt evolution (differentiation)—temperature relationships. There is still a region (II, Fig. 14) that is dominated by rapid decrease in melt fraction and enhanced differentiation, but both effects are much weaker for fractional than for the equilibrium crystallization. As pointed out by Melekhova et al. (2013), the strongly nonlinear crystallization behavior with respect to temperature dictates the chemical evolution of magmas differentiating in mid-crustal magma reservoirs and it most likely also controls the dynamics of magma reservoirs and the extraction of liquids from solid residues: At the onset of crystallization, minor olivine and clinopyroxene crystallization generates phenocryst-poor magmas that form dilute suspensions; magma viscosity remains low, close to the liquid viscosity (e.g., Caricchi et al. 2007; Costa et al. 2009; Pistone et al. 2012). Separation of crystals and liquid is difficult in particular when magmas are convecting or rapidly ascending. The significant increase in crystallinity over a small temperature interval (II, Fig. 14) brings the magma close to or within the rigid percolation threshold (RPT; Vigneresse et al. 1996) corresponding to the formation of a randomly loose packed crystal framework and coincides with the transition from a dilute to a concentrated suspension (Pistone et al. 2013). Such rheological conditions enable the partitioning of strain to form melt-enriched (dilatant) regions and crystal-rich (compressive) regions ultimately generating pathways to efficiently extract liquids from solids in a deforming crystal mush. Thus, we propose that the rapid increase in the crystal content in the crystallizing magma over a short temperature interval generates ideal rheological conditions for melt extraction just after the switch from pyroxene to amphibole-dominated solid



**Fig. 14** Melt fraction relative to initial mass of melt and solid fraction relative to initial mass of melt (wt%) as a function of temperature. For comparison and discussion, the melt fractions obtained by Melekhova et al. (2013) for equilibrium crystallization of a primitive Mg-rich basalt at 0.7 GPa for 2.3 and 4.5 wt% H<sub>2</sub>O and by Blatter et al. (2013) at 0.7 GPa for a basalt with 2.0 wt% H<sub>2</sub>O are indicated

assemblages. The extracted liquids would be a typical andesite, the most common bulk rock composition observed in plutonic and volcanic rocks in convergent plate margin magmatism (e.g., Gill 1981). Therefore, the dominance of andesitic compositions in supra-subduction igneous provinces is probably just the consequence of the topology of the underlying phase diagram and the thermodynamically controlled phase occurrence and proportions. Under mid-crustal conditions investigated in this study, basaltic liquids differentiate over a small temperature interval (60–80 °C) to andesite and produce an amount of solid phases that changes the rheology of the suspension to enable extraction of the liquid from the crystal mush. The buoyant andesitic magmas extracted from the crystal mush most likely ascend further through the crust and lead to the formation of plutonic complexes in shallower crustal levels. Provided governing phase assemblages and modal proportions do not change significantly at somewhat lower pressure, the present experiments imply that this andesitic magma will not reach the RPT until approaching the granite minimum at temperatures around 710 °C when a total amount of 50 % of the extracted andesite magma will have crystallized to amphibole + plagioclase + magnetite (reducing the melt fraction from 40 to 20 % relative to the starting basalt). Over this large temperature interval (280 °C), the magma potentially behaves as a dilute suspension, which not only allows rapid ascent in dikes, but also further enables the buildup of highly dynamic, connecting magma reservoirs within the crust. Only when the granite minimum is approached, will the magma reach the RPT where the extraction of a high-silica-rich rhyolitic liquid could occur (e.g., Bachmann and Bergantz 2004).

The scenario depicted in the previous paragraph cannot be strictly derived from the experimental data set presented in this study for the following reasons: (1) the processes envisaged propose polybaric evolution from a lower mid-crustal magma reservoir generating the andesitic liquids that is emplaced at shallower level forming intermediate to acidic plutonic complexes or even erupt at the surface of the Earth. This would require polybaric fractional crystallization experiments and/or a coherent data set to predict polybaric behavior; and (2) the scenario implicitly does not represent (near) perfect fractional crystallization but a combination of fractional and equilibrium crystallization; such a concept is also favored by Melekhova et al. (2013). Liquids are only extracted when the rheologic RPT is reached; thus, magmas are allowed to equilibrate (at least to a certain extent) with the solid residue. However, the amphibole + plagioclase + magnetite crystallization interval (III) is characterized by cotectic crystallization of the observed phases, and thus, we do not anticipate significant difference between fractional and equilibrium crystallization regarding melt compositions and melt fractions potentially allowing to apply the results of this study directly to natural systems. Nevertheless, equilibrium crystallization experiments that partly exist (Kawamoto 1996; Sisson et al. 2005; Almeev et al. 2013; Blatter et al. 2013; Melekhova et al. 2013) should be combined with this study to better constrain phase assemblages, liquid compositions and melt fractions as a function of temperature and pressure with the ultimate goal to provide a consistent model of convergent margin magmatism that not only takes into account thermodynamic (chemical and energetic) constraints but also physical (rheologic, density) considerations to understand the crustal evolution of such magmatic systems.

## Conclusion

Near-perfect fractional crystallization experiment at 0.7 GPa from the hydrous basalt liquidus close to the H<sub>2</sub>O-saturated granite solidus results that the presence of moderate amounts of H<sub>2</sub>O (3 wt%) in the parental magma is sufficient to drive the liquid line of descent from tholeiitic (anhydrous) to calc-alkaline (hydrous) by suppressing crystallization of plagioclase and enhancing the crystallization of clinopyroxene and magnetite thereby inhibiting iron enrichment characterizing the tholeiitic evolutionary trend. Silica enrichment, the other prominent characteristic of calc-alkaline magmas, is achieved by crystallization of anorthitic plagioclase, amphibole and Fe–Ti oxide. Amphibole stability in intermediate calc-alkaline magmas is governed by a diverging peritectic encompassing the amphibole stability field. Depending on pressure and



probably also H<sub>2</sub>O and Na<sub>2</sub>O contents of the system, the amphibole stability field is delimited by reactions (1) involving clinopyroxene + liquid ± plagioclase and (2) olivine or orthopyroxene + liquid ± plagioclase. The amphibole stability field is known to be rather small in compositional space (Helz 1976; Allen and Boettcher 1983; Huang and Wyllie 1986; Hilyard et al. 2000) and, therefore, sensitive to small changes in the solid phase assemblages crystallized before reaching the amphibole field. In the present case, the system crystallized along the orthopyroxene–amphibole–plagioclase reaction boundary (Fig. 12) to peraluminous derivative liquids. In order to reach the amphibole stability field, extensive fractionation of clinopyroxene accompanied by minor plagioclase and spinel extraction, solidifying 50 % of the original melt mass within a short temperature interval must take place; this might also occur at lower levels within the crust or even within the uppermost mantle. Amphibole systematically varies in composition with decreasing temperature from nepheline- to quartz-normative and significantly contributes to drive derivative liquids to peraluminous compositions. A typical gabbroic mineral assemblage composed of amphibole, plagioclase and magnetite crystallizes over a 280 °C temperature interval while coexisting liquids differentiate from andesite to high-silica rhyolite. The formation of high-silica rhyolite from andesite at mid-crustal levels requires about 50 % extraction of gabbroic cumulates leaving about 20 % rhyolite relative to the mass of the initial basaltic mantle-derived magma.

Corundum-normative (Chayes 1969) volcanics and peraluminous granitoid rocks (Williams et al. 1955) are a common product in supra-subduction zone igneous systems. This experimental study demonstrates that such compositions can be produced from primitive mantle-derived basaltic melts by fractional crystallization at 0.7 GPa corresponding to mid-crustal levels.

The strongly nonlinear melt fraction–temperature relationships along the liquid line of descent potentially results in two temperature-compositional windows where melt extraction is most effective; a first one just above 1,000 °C where an andesitic liquid coexists with an ultramafic (clinopyroxenitic) to mafic (gabbro to amph–gabbro) crystal mush and a second one below 710 °C where a high-silica rhyolite coexists with an amphibole-bearing diorite/tonalite. Above 1,070 °C and between 1,000 and 710 °C, crystal poor to moderately crystalline basalts, basaltic andesites and andesites form highly mobile melts potentially ascending through the crust reaching higher-level reservoirs.

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